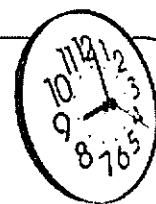


# Model Test Paper-4



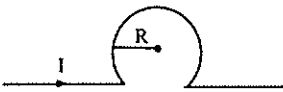
Time :  $3\frac{1}{2}$  hours.

Maximum Marks : 200

## PHYSICS

- In Huygen's wave theory, the focus of all points in the same state of vibration called
  - wave front
  - a half-period zone
  - vibrator
  - a ray
- A spring has been kept fixed with immovable wall and a force of 1 newton has been applied on it. The spring extends upto a length  $l$ . Now if 2 newton force is applied on it, it will extend
  - $\frac{2}{3} l$
  - $\frac{3}{2} l$
  - $2l$
  - $l^2$
- Refractive index is greater for
  - light of greater wavelength
  - light of shorter wavelength
  - light of low frequency
  - all of these
- A thin circular copper plate, a sphere and a cube of same mass and material are heated to  $100^\circ\text{C}$  temperature. Now if they are allowed to cool, which of the three will cool first ?
  - sphere
  - copper plate
  - cube
  - all at same time
- If we bring N-pole of a magnet towards the coil. The face of the coil acquires
  - zero-polarity
  - north polarity
  - south polarity
  - none of these
- In a closed and organ pipe which of the following notes is not present if fundamental note is 50?
  - 100
  - 250
  - 150
  - none of these
- In a semiconductor diode P-side is earthed and N-side is applied a potential of 2 volt, the diode shall
  - breakdown
  - not conduct
  - conduct
  - conduct partially
- The sensitivity of galvanometer depends on
  - moment of inertia of coil
  - angle of deflection
  - earth's magnetic field
  - none of these
- If  $v_e$  be the escape velocity and  $v_o$  be the orbital velocity, then  $v_e/v_o$  is equal to
  - $2\sqrt{2}$
  - $\sqrt{2}$
  - $\frac{1}{\sqrt{2}}$
  - 2
- A pendulum of time period  $T$  is kept suspended in a train accelerating uniformly, then its time period
  - decreases
  - increases
  - remains unchanged
  - none of these
- The difference in the acceleration due to gravity at the pole and equator is given by
  - $R\omega^2$
  - $\Omega \cos \theta^2$
  - $R\omega^2 \cos^2\theta$
  - $\frac{R\omega^2\theta}{g^2}$
- If two bulbs one of 60 W and other of 100 W are connected in parallel, then which one of the following will glow more ?
  - 60 W bulb
  - 100 W bulb
  - both equally
  - none of these
- The number of electrons ejected from photoelectric surface depends upon
  - the wavelength of light
  - the frequency of light
  - the intensity of incident light
  - none of these

14. An  $n$ -type semiconductor is formed  
 (a) only from germanium  
 (b) when germanium is doped with impurity containing  $3d$  valence electrons  
 (c) when germanium is doped with impurity containing  $5$ -valence electrons  
 (d) only from pure silicon
15. Which one of the following is essential feature of SHM ?  
 (a) acceleration is directly proportional to displacement from mean position and is directed towards it  
 (b) restoring force is inversely proportional to displacement from mean position  
 (c) acceleration and amplitude  
 (d) constant amplitude
16. Which of the following wavelengths will suffer maximum deviation while passing through a prism?  
 (a) orange (b) green  
 (c) violet (d) red
17. A spectrum which contains all wavelengths without any break is called  
 (a) continuous emission spectrum  
 (b) line spectrum  
 (c) emission spectrum  
 (d) all of these
18. Seebeck emf depends on  
 (a) neutral temperature  
 (b) temperature of cold junction  
 (c) temperature of hot junction  
 (d) none of these
19. A hollow cylinder and a solid cylinder having same mass and same diameter are released from rest simultaneously from top of an inclined plane which one will reach bottom first ?  
 (a) solid (b) hollow  
 (c) both equally  
 (d) one with greater density
20. A hollow charged metal sphere has radius  $r$ . If the potential difference between its surface and a point at a distance of  $3r$  from the centre is  $V$ , then electric field intensity at a distance of  $3r$  from centre is  
 (a)  $\frac{V}{2}r$  (b)  $\frac{V}{4}r$   
 (c)  $-\frac{V}{3r}$  (d)  $\frac{V}{6}r$
21. In an A.C. circuit  $V$  and  $I$  is given by  
 $V = 100 \sin(1000t)$  volt  
 $I = 1000 \sin(1000t + \pi/3)$  mA The power dissipated in the circuit is  
 (a) 10 W (b) 25 W  
 (c)  $10^4$  W (d) 250 W
22. A series LCR circuit is tuned to resonance. The angular frequency of the applied AC voltage is  $\omega$ . If resistance of the circuit is  $R$ , the impedance of circuit will be  
 (a)  $R^2 + \left(\omega L + \frac{1}{\omega C}\right)^2$  (b)  $R + \omega L + \left(\frac{1}{\omega C}\right)$   
 (c)  $R$  (d)  $\sqrt{R^2 + \left(\omega L - \frac{1}{\omega C}\right)^2}$
23. Choose the correct answer  
 Heat in metals is produced due to  
 (a) collision of conduction electrons with protons  
 (b) collision of conduction electrons with atoms  
 (c) collision of electrons with electrons  
 (d) in all these ways as mentioned in above options
24. A proton enters a magnetic field parallel to the direction of field, then the path following by it is  
 (a) straight line (b) hyperbola  
 (c) circular (d) helical
25. The negative sign in the equation  
 $e = \frac{-d\phi}{dt}$  indicates  
 (a) induced emf opposes the cause producing it  
 (b) current density is negative  
 (c) emf is always taken negative  
 (d) none of the above

26. Fission reaction was discovered by  
 (a) Seaborg  
 (b) Otto Han and Strassman  
 (c) Einstein (d) S. Hawking
27. If 200 MeV energy is released in a fission of a single nucleus of  ${}_{92}\text{U}^{235}$ . How many fissions must occur per second to produce a power of kW ?  
 (a)  $3.125 \times 10^{13}$  (b)  $3.12 \times 10^{12}$   
 (c)  $0.312 \times 10^{13}$  (d) none of these
28. Some water drops of radius  $r$  each coalesce to form a big drop of radius  $R$ . Then rise in temperature is given by  
 (a)  $\left(\frac{3T}{J}\right)\left(\frac{1}{r}-\frac{1}{R}\right)$  (b)  $\frac{3T}{J.r}$   
 (c)  $\frac{rT}{J}$  (d)  $\frac{3T}{J}\left(\frac{1}{r}+\frac{1}{R}\right)$
29. A long straight conductor is bent into shape as shown. If it carries  $I$  ampere and its radius is  $R$ , then magnetic field ( $\vec{B}$ ) at the centre of circular coil is  
  
 (a)  $\frac{\mu_0 I(\pi + 1)}{2\pi R}$  (b) Zero  
 (c)  $\infty$  (d)  $\frac{\mu_0 I(\pi - 1)}{2\pi R}$
30. An engineer claims to have made an engine delivering 10 kW power with fuel consumption of  $1\text{g sec}^{-1}$ . The calorific value of fuel is  $2\text{k cal/gm}$ . His claim is  
 (a) depends on engine  
 (b) valid (c) non-valid  
 (d) depends on load
31. If two drops of same radius are falling through air with a velocity of  $5\text{ cm sec}^{-1}$ . If the two drops coalesce to form one drop, the terminal velocity of the drop is  
 (a)  $5\sqrt{2}\text{ cm/sec}$  (b)  $10\text{ cm/sec}$   
 (c)  $2.5\text{ cm/sec}$  (d)  $5 \times 4^{1/3}\text{ cm/sec}$
32. The frequency of open organ pipe is  $f$ . If half part of the organ pipe is dipped in water, then frequency is  
 (a)  $3f/4$  (b)  $f/2$   
 (c)  $f$  (d) zero
33. A particle is acted upon by a force of constant magnitude which is always perpendicular to the velocity of particle. The motion of the particle takes place in a horizontal plane. It follows  
 (a) it moves in a circular path  
 (b) velocity is constant  
 (c) linear momentum is constant  
 (d) none of these
34. The number of turns in the coil are doubled, the emf will get  
 (a) quadrupled (b) halved  
 (c) doubled (d) none of these
35. Nucleus contains  
 (a) protons, electrons and neutrons  
 (b) electrons and neutrons  
 (c) protons and electrons  
 (d) protons and neutrons
36. Diamond shines due to  
 (a) total internal reflection  
 (b) refraction (c) reflection  
 (d) none of these
37. The frequency of sonometer wire is  $n$ . If its tension is increased four times and length is doubled, the new frequency will be  
 (a)  $n$  (b)  $2n$   
 (c)  $n/2$  (d)  $4n$
38. Which of the following combination would give maximum emf ?  
 (a) Sb and Bi (b) Fe and Bi  
 (c) Ni and Cr (d) Cu and Fe
39. According to Bohr's theory, the radius of electron in an orbit described by principal quantum number  $n$  and atomic number  $Z$  is proportional to  
 (a)  $\frac{n^2}{Z^2}$  (b)  $\frac{n^2}{Z}$   
 (c)  $Z^2 n^2$  (d)  $\frac{Z^2}{n}$

40.  $\frac{\text{kg.m}^2}{\text{s}^2}$  is the unit of

- (a) Momentum                      (b) Power  
(c) Energy                            (d) Impulse

**Instructions for Q. No. 41 to 60**

*Directions : Each of the questions given below consists of two statements, an assertions (A) and reason (R). Select the number corresponding to the appropriate response in the answer sheet as follows.*

- (a) If both assertion and reason are true and the reason is a correct explanation of the assertion  
(b) If both assertion and reason are true but the reason is not a correct explanation of the assertion  
(c) If the assertion is true, but the reason is false  
(d) If both assertion and reason are false
41. *Assertion (A) :* A dip needle becomes vertical at magnetic equator of the earth.  
*Reason (R) :* The magnetic field due to the earth at the magnetic equator is vertical.
42. *Assertion (A) :* When two electrons are brought close to each other, the electrical potential energy increases.  
*Reason (R) :* Work must be done against electrical force of repulsion.
43. *Assertion (A) :* If Young's double slit experiment is performed in water, the fringe width will decrease  
*Reason (R) :* Wavelength of light in water is smaller than in air.
44. *Assertion (A) :* Interference pattern is obtained on a screen due to two identical coherent sources of monochromatic light. The intensity at the central part of the screen becomes one half if one of the sources is blocked.  
*Reason (R) :* The resultant density is the sum of the densities due to two sources; if one is blocked the intensity obviously reduces to one-half.
45. *Assertion (A) :* Insulators do not allow flow of current through them.  
*Reason (R) :* They have no free charge carriers.
46. *Assertion (A) :* The shape of an automobile is so designed that its front resembles the streamline pattern of the fluid through which it moves.  
*Reason (R) :* the resistance offered by the fluid is maximum.
47. *Assertion (A) :* Two satellites of mass  $m_1$  and  $m_2$  ( $m_1 > m_2$ ) are going around the earth in orbits of radii  $r_1$  and  $r_2$  ( $r_1 > r_2$ ).  
*Reason (R) :* They will have same velocity.
48. *Assertion (A) :* In Thomson's experiment all the positive ions with the same value of specific charge are focussed on the same parabola irrespective of their velocities.  
*Reason (R) :* The ions of same velocities arrive at different points on the same parabola.
49. *Assertion (A) :* In the process of nuclear fission the fragments emit two or three neutrons as soon as they are formed and subsequently emit particles.  
*Reason (R) :* As the fragments contain an excess of neutrons over protons emission of neutrons and particles bring their neutron/proton ratio to stable values.
50. *Assertion (A) :* While passing round the corners of an obstacle the light spreads out to some extent into the region of the geometrical shadow.  
*Reason (R) :* The bending is greater for light of longer wavelengths and less for shorter wavelengths.
51. *Assertion (A) :* An e.m.f. is induced in a circuit whenever there is a change in the magnetic flux linked with the circuit and the magnitude of the induced e.m.f. is equal to the negative rate of change of flux.  
*Reason (R) :* The direction of the induced e.m.f. is such that it opposes the very cause to which it is due.
52. *Assertion (A) :* If a heavy nucleus is split into two medium sized parts, each of the new nuclei will have more binding energy per-nucleon than the original nucleus.  
*Reason (R) :* Combining two light nuclei to form a single relatively heavy nucleus means more binding energy per nucleon in the new nucleus.

53. *Assertion (A)* : When two vibrating tuning forks having frequencies 256Hz and 512Hz are held near each other, beats can not be heard.  
*Reason (R)* : The principle of superposition is valid only if the frequencies of the oscillators are nearly equal.
54. *Assertion (A)* : In the absence of space charge, the potential gradient between cathode and the anode will be uniform.  
*Reason (R)* : The space charge reduces the potential in the cathode and anode region non-uniformly.
55. *Assertion (A)* : A thin aluminium disc, spinning freely about a central pivot, is quickly brought to rest when placed between the poles of strong U-shaped magnet.  
*Reason (R)* : Current induced in the disc rotating in a magnetic field produces a force which opposes the motion of the disc.
56. *Assertion (A)* : When white light is incident on a thin oil film on the surface of water, colours are seen.  
*Reason (R)* : White light is composed of several colours.
57. *Assertion (A)* : The work done in bringing a body from the top to the base along a frictionless inclined plane is the same as the work done in bringing it down along the vertical side.  
*Reason (R)* : The gravitational force on the body along the inclined plane is the same as that along the vertical side.
58. *Assertion (A)* : A vibrating tuning fork sounds louder when its stem is put against a desk top.  
*Reason (R)* : When a wave reaches another denser medium, a part of the wave is reflected.
59. *Assertion (A)* : Isotopes of an element can be separated by using a mass spectrometer.  
*Reason (R)* : Separation of isotopes is possible because of the difference in electron numbers of isotopes.
60. *Assertion (A)* : A large soap bubble expands while a small bubble shrinks, when they are connected to each other by a capillary tube.  
*Reason (R)* : The excess pressure inside bubble (or a drop) is inversely proportional to its radius

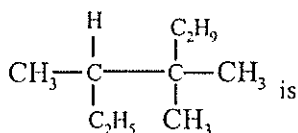
## CHEMISTRY

61. Consider the following reaction occurring in an automobile  

$$2C_8H_{18(g)} + 25O_{2(g)} \rightarrow 16CO_{2(g)} + 18H_2O_{(g)}$$
 The sign  $\Delta H$ ,  $\Delta S$  and  $\Delta G$  would be  
 (a) -, +, + (b) -, +, -  
 (c) +, -, + (d) +, +, -
62. Equivalent conductance of NaCl, HCl and  $CH_3COONa$  at infinite dilution are 126.45, 426.16 and  $91 \text{ ohm}^{-1} \text{ cm}^2$  respectively. The equivalent conductance of  $CH_3COOH$  at infinite dilution would be  
 (a)  $390.71 \text{ ohm}^{-1} \text{ cm}^2$  (b)  $253.71 \text{ ohm}^{-1} \text{ cm}^2$   
 (c)  $101.38 \text{ ohm}^{-1} \text{ cm}^2$  (d)  $678.90 \text{ ohm}^{-1} \text{ cm}^2$
63. Which does not exist ?  
 (a)  $[CCl_6]^{2-}$  (b)  $[GeF_6]^{2-}$   
 (c)  $[SiCl_6]^{2-}$  (d)  $[GeF_6]^{2-}$
64. The type of hybridisation of boron in diborane is  
 (a)  $sp^3$  (b)  $sp^2$   
 (c)  $sp^2$  (d)  $sp^3d^2$
65. Boron compounds behave as Lewis acids because of their  
 (a) electron deficient character  
 (b) covalent nature  
 (c) acidic nature  
 (d) ionising property
66. All the following substances react with water. The pair that yields the same gaseous product is  
 (a) Ca and  $CaH_2$  (b) Na and  $Na_2O_2$   
 (c) K and  $KO_2$  (d) Ba and  $BaO_2$
67. Epsom salt is  
 (a)  $2CaSO_4 \cdot H_2O$  (b)  $MgSO_4 \cdot 2H_2O$   
 (c)  $MgSO_4 \cdot 7H_2O$  (d)  $BaSO_4 \cdot 2H_2O$
68. Squashes are stored by adding  
 (a)  $Na_2SO_3$  (b) KCl  
 (c) Citric acid  
 (d) Sod. metabisulphite
69. In Goldschmidt aluminothermic process, reducing agent used is  
 (a) Na (b) Al powder  
 (c) Coke (d)  $Al_2O_3$

70. Which one of the following pairs of substances on reaction will not evolve  $H_2$  gas ?  
 (a) iron and  $H_2SO_4$  (aq)  
 (b) iron and steam  
 (c) copper and  $HCl$  (aq)  
 (d) sodium and ethanol
71. The oxide that gives hydrogen peroxide on treatment with dil acid is  
 (a)  $MnO_2$  (b)  $Na_2O_2$   
 (c)  $PbO_2$  (d)  $TiO_2$
72. The energy required to remove an electron of a gaseous atom from its ground state is called  
 (a) electrode potential  
 (b) ionisation energy  
 (c) potential energy  
 (d) activation energy
73. Chloride of an element A gives neutral solution in water. In the periodic table, the element A belongs to  
 (a) fifth group  
 (b) third group  
 (c) first group  
 (d) first transition series
74. If the valency shell electronic configuration of an element is  $ns^2 np^5$ , this element belongs to the group of  
 (a) noble gases (b) inert metals  
 (c) alkali metals (d) halogens
75. Main product of reaction  $CH_3CONH_2 + HNO_2 \rightarrow ?$  is  
 (a)  $CH_3NH_2$  (b)  $CH_3CH_2NH_2$   
 (c)  $CH_3COOH$  (d)  $CH_3NO_2$
76. Paraldehyde is  
 (a) a hexamer of formaldehyde  
 (b) a trimer of acetaldehyde  
 (c) a trimer of formaldehyde  
 (d) a hexamer of acetaldehyde
77. An aldehyde when treated with an alkali gives an acid and an alcohol. Such reaction is named as  
 (a) cannizzaro reaction  
 (b) hydrolysis reaction  
 (c) aldol condensation  
 (d) claisen condensation
78. Reaction  $CO + H_2 + H_2 \xrightarrow[Cr_2O_3-ZnO]{673 K, 300 atm}$  may be used for the manufacture of  
 (a)  $HCHO$  (b)  $HCOOH$   
 (c)  $CH_3OH$  (d)  $CH_3COOH$
79. An organic compound A reacts with sodium metal and forms B. On heating with conc.  $H_2SO_4$ , A gives diethyl ether. A and B are respectively  
 (a)  $CH_3OH, CH_3ONa$   
 (b)  $C_3H_7OH, C_3H_7ONa$   
 (c)  $C_2H_5OH, C_2H_5ONa$   
 (d)  $C_4H_9OH, C_4H_9ONa$
80. Glucose molecule reacts with 'X' number of phenyl hydrazine molecules to yield osazone. The value of 'X' is  
 (a) 3 (b) 2  
 (c) 1 (d) 4
81. The sugar present in fruits is  
 (a) sucrose (b) glucose  
 (c) fructose (d) galactose
82. To become a carbohydrate a compound must contain at least  
 (a) 4 carbons (b) 3 carbons  
 (c) 2 carbons (d) 6 carbons
83. Vitamin D is also known as  
 (a) reproductive vitamin  
 (b) ascorbic acid  
 (c) growth vitamin  
 (d) sunshine vitamin
84. Zwitter ion is formed by  
 (a) benzoic acid (b) acetanilide  
 (c) aniline (d) lysine
85. Aspirin is an acetylation product of  
 (a) m-hydroxybenzoic acid  
 (b) o-dihydroxy benzene  
 (c) o-hydroxybenzoic acid  
 (d) p-dihydroxy benzene
86. The olefin which on ozonolysis gives  $CH_3CH_2CHO$  and  $CH_3CHO$  is  
 (a) 1-pentene (b) 2-butene  
 (c) 1-butene (d) 2-pentene
87. An isomer of ethanol is  
 (a) diethyl ether (b) dimethyl ether  
 (c) methanol (d) ethylene glycol

88. IUPAC name of



- (a) 3, 4, 4-trimethyl heptane  
 (b) 2-ethyl-3, 3-dimethyl heptane  
 (c) 2-butyl-2-methyl-3 ethyl butane  
 (d) 3, 4, 4-trimethyl octane
89. The most-suitable method of separation of 1 : 1 mixture of *o*- and *p*-nitrophenols is  
 (a) crystallisation (b) chromatography  
 (c) sublimation (d) steam distillation
90. Consider the reaction  

$$\text{M}^{n+}_{(\text{aq})} + n\text{e}^- \rightarrow \text{M}_{(\text{s})}$$
  
 The standard reduction potential values of the metals  $\text{M}_1$ ,  $\text{M}_2$  and  $\text{M}_3$  are  $-0.34\text{V}$ ,  $-33.05\text{V}$  and  $-1.66\text{V}$  respectively. The order of their reducing power will be  
 (a)  $\text{M}_1 > \text{M}_3 > \text{M}_2$  (b)  $\text{M}_3 > \text{M}_2 > \text{M}_1$   
 (c)  $\text{M}_1 > \text{M}_2 > \text{M}_3$  (d)  $\text{M}_2 > \text{M}_3 > \text{M}_1$
91. Specific conductance of 0.1 N KCl solution at  $25^\circ\text{C}$  is  $0.012 \text{ ohm}^{-1} \text{ cm}^{-1}$ . The resistance of the cell containing the solution at the same temperature was found to be 55 ohm. The cell constant will be  
 (a)  $0.918 \text{ cm}^{-1}$  (b)  $0.66 \text{ cm}^{-1}$   
 (c)  $0.142 \text{ cm}^{-1}$  (d)  $1.12 \text{ cm}^{-1}$
92.  $\text{S} + \frac{3}{2} \text{O}_2 \rightarrow 3\text{O}_3 + 2x \text{ k cal}$   

$$\text{SO}_2 + \frac{1}{2} \text{O}_2 \rightarrow \text{SO}_3 + y \text{ k cal}$$
  
 Find out the heat of formation of  $\text{SO}_2$   
 (a)  $(x + y)$  (b)  $(2x + y)$   
 (c)  $(2x - y)$  (d)  $\frac{2x}{y}$
93. How many layers are adsorbed in chemical adsorption?  
 (a) several (b) 2  
 (c) 1 (d) zero
94. Activation energy of a chemical reaction can be determined by  
 (a) evaluating rate constants at two different temperatures  
 (b) evaluating velocities of reaction at two different temperatures  
 (c) evaluating rate constant at standard temperature  
 (d) changing concentration of reactants
95. The pH of a solution obtained by mixing 50 ml 0.4N HCl and 50 ml 0.2N NaOH is  
 (a) 1.0 (b)  $-\log 0.2$   
 (c)  $-\log 2$  (d) 2.0
96. In which case  $K_p$  is less than  $K_c$ ?  
 (a)  $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$   
 (b)  $\text{H}_2 + \text{Cl}_2 \rightleftharpoons 2\text{HCl}$   
 (c)  $\text{PCl}_5 \rightleftharpoons \text{PCl}_3 + \text{Cl}_2$   
 (d)  $2\text{SO}_3 + \text{O}_2 \rightleftharpoons 2\text{SO}_5$
97. How many grams of dibasic acid (mol. wt. 200) should be present in 100 ml of the aqueous solution to give 0.1N normality?  
 (a) 2 g (b) 20 g  
 (c) 1 g (d) 10 g
98. The ratio between the two mean square speed of  $\text{H}_2$  at 50 K and that of  $\text{O}_2$  at 800 K is  
 (a) 1 (b) 2  
 (c) 4 (d)  $\frac{1}{4}$
99. If we mix a pentavalent impurity in a crystal lattice of germanium, what type of semi-conductor formation will occur?  
 (a) *p*-type (b) *n*-type  
 (c) both (a) and (b) (d) none of the two
100. A solid has a structure in which 'W' atoms are located at the corners of a cubic lattice, 'O' atoms at the centre of edges and 'Na' atoms at the centre of the cube. The formula of the compound is  
 (a)  $\text{Na}_2\text{WO}_3$  (b)  $\text{NaWO}_3$   
 (c)  $\text{NaWO}_2$  (d)  $\text{NaWO}_4$

**Instructions for Q. No. 101 to 120**

*Directions : Each of the questions given below consists of two statements, an assertions (A) and reason (R). Select the number corresponding to the appropriate response in the answer sheet as follows.*

- (a) *If both assertion and reason are true and the reason is a correct explanation of the assertion*
- (b) *If both assertion and reason are true but the reason is not a correct explanation of the assertion*
- (c) *If the assertion is true, but the reason is false*
- (d) *If both assertion and reason are false*
101. *Assertion (A) : Lead is a metal with a high density. It readily dissolves in moderately concentrated nitric acid giving colourless fumes which turn red in contact with air.*  
*Reason (R) : Nitric oxide (NO) is a colourless oxide of nitrogen while NO<sub>2</sub> is a coloured oxide of nitrogen.*
102. *Assertion (A) : The reaction of ammonia solution with calomel is a disproportionation reaction in which mixture of Hg (ii) amido chloride and Hg are formed.*  
*Reason (R) : In a disproportionation reaction species under reaction is neither oxidised nor reduced.*
103. *Assertion (A) : Sodium thiosulphate dissolves the white precipitate of silver chloride.*  
*Reason (R) : The thiosulphate ions act as strong complexing agents.*
104. *Assertion (A) : When SnCl<sub>2</sub> solution is added to HgCl<sub>2</sub> solution, a milky white precipitate is obtained and on adding excess of SnCl<sub>2</sub>, a black precipitate is formed.*  
*Reason (R) : The disproportionation of Hg(II) is easier than its reduction only.*
105. *Assertion (A) : The electron affinity of chlorine is greater than that of fluorine.*  
*Reason (R) : Chlorine is more electronegative than fluorine.*
106. *Assertion (A) : The boiling point of n-alkanes increases regularly with the increase in the number of carbon atoms.*  
*Reason (R) : The magnitude of van der Waal's forces increases with the increases in molecular mass and molecular size.*
107. *Assertion (A) : p-nitroaniline is stronger base than p-toluidine.*  
*Reason (R) : The electron withdrawing NO<sub>2</sub> group in the p-nitroaniline makes it a stronger base.*
108. *Assertion (A) : All the amines, except tertiary amines are capable of forming intermolecular hydrogen bonds.*  
*Reason (R) : Tertiary amines have larger molecules and surface area.*
109. *Assertion (A) : Phenol is strongly acidic than ethanol.*  
*Reason (R) : Phenoxide ion is more stabilized by resonance than ethoxide ion.*
110. *Assertion (A) : The nuclear isomers are the atoms with the same atomic number and same mass number, but with different radioactive properties.*  
*Reason (R) : The nucleus in the excited state will evidently have a different half-life as compared to that in the ground state.*
111. *Assertion (A) : Balloons made of Mylar films are better at containing helium than the conventional rubber balloons*  
*Reason (R) : The root-mean-square speed of helium is very high so helium atoms can effuse rapidly through rubber balloons.*
112. *Assertion (A) : To separate <sup>235</sup>U from the more abundant <sup>238</sup>U isotope, all the uranium is converted into UF<sub>6</sub>.*  
*Reason (R) : UF<sub>6</sub> is one of the few compounds that exists in gaseous state under ordinary conditions.*
113. *Assertion (A) : One mole of helium atoms should occupy 22.4 litre volume at STP.*  
*Reason (R) : Taking 31 pm as radius of helium atom, if we pack together a mole of helium atoms, the mole of atoms should have a volume of 22.4 litre.*



114. *Assertion (A)* : A sample of 8.00 moles of chlorine gas in a 4.00 litre tank 27 C leads to a pressure of 49.2 atm according to ideal gas law.  
*Reason (R)* : The actual pressure of the sample of chlorine is nearly 20 atmosphere less than the ideal pressure.
115. *Assertion (A)* : The pressure of a gas is inversely proportional to its volume at constant temperature and  $n$ .  
*Reason (R)* : The gas volume is directly proportional to  $n$  at constant temperature and pressure.
116. *Assertion (A)* : Not only is the fraction of oxygen is reduced in diving gases, but nitrogen of normal air is replaced by helium.  
*Reason (R)* : Nitrogen becomes more soluble in the body fluids at high pressure and causes a condition similar to alcohol intoxication.
117. *Assertion (A)* : When one talks after breathing helium, the sound becomes like that of Donald Duck.  
*Reason (R)* : The vocal cords vibrate faster in an atmosphere less dense than air and the pitch of voice is raised.
118. *Assertion (A)* : The reacting gases combine in volumes that are ratios of small whole numbers.  
*Reason (R)* : The partial pressure of a gas in a mixture is given by its mole fraction times the total pressure of the mixture.
119. *Assertion (A)* : The oxidation numbers are artificial, they are useful as a 'book-keeping' device of electrons in reactions.  
*Reason (R)* : The oxidation numbers do not usually represent real charges on atoms, they are simply conventions that indicate what the maximum charge could possibly be on an atom in a molecule.
120. *Assertion (A)* : The structural-pair geometry of Formaldehyde molecule is trigonal planar.  
*Reason (R)* : In  $H_2CO$  molecule, the carbon atom is surrounded by 3 sigma bonding electron pairs.

## BIOLOGY

121. In which of the following animal, all the three important chordate characters exist throughout life ?  
 (a) amphibians (b) mammals  
 (c) *Amphioxus* (d) all of the above
122. Which cranial nerve has the highest number of branches ?  
 (a) vagus nerve (b) facial nerve  
 (c) trigeminal (d) all of the above
123. What is common among silver fish, crab, honey bee and prawn ?  
 (a) metamorphosis (b) compound eye  
 (c) poison gland (d) all of the above
124. The maximum formation of m-RNA occurs in  
 (a) ribosome (b) nucleolus  
 (c) cytoplasm (d) nucleoplasm
125. The most striking example of point mutation is found in a disease, called  
 (a) down's syndrome (b) night blindness  
 (c) thalassemia (d) sickle-cell anaemia
126. At high altitude, the RBCs in the human blood will  
 (a) increase in number  
 (b) decrease in size  
 (c) increase in size (d) decrease in number
127. Typhus disease in humans is caused by  
 (a) rickettsiae (b) protozoans  
 (c) virus (d) none of the above
128. Rickettsiae form a group of  
 (a) bacterium-like prokaryotes  
 (b) viruses (c) fungi  
 (d) none of the above
129. In the fertile human female, approximately on which day of the ovulation takes place ?  
 (a) 14th day (b) 8th day  
 (c) 1st day (d) 18th day
130. Which of the following is regarded as an unit of nervous tissue?  
 (a) neuron (b) dendrite  
 (c) axon (d) myelin sheath

131. Which of the following carries absorbed product from digestive tract ?  
(a) pulmonary vein  
(b) hepatic portal vein  
(c) hepatic artery  
(d) none of the above
132. Who proposed the 'signal hypothesis' meant for the biosynthesis of secretory type of proteins?  
(a) Blobel and Sabatini  
(b) Camillo Golgi (c) Baltimore  
(c) Sheeler and Bianchi
133. Which of the following carries protein and lipid to other parts of the cell?  
(a) rough endoplasmic reticulum  
(b) smooth endoplasmic reticulum  
(c) both (a) and (b)  
(d) none of the above
134. Epidermal layer consisting of dividing cells, is  
(a) stratum granulosum  
(b) stratum malpighii  
(c) stratum lucidum  
(d) stratum corneum
135. The tissue having least power of regeneration is  
(a) skeletal tissue of long bones  
(b) endothelium of blood vessels  
(c) epidermis of skin  
(d) nervous tissue of brain
136. Which of the following evidences does not favour the Lamarckian concept of inheritance of acquired characters ?  
(a) lack of pigment in cave dwelling animals  
(b) presence of webbed toes in aquatic birds  
(c) absence of limbs in snakes  
(d) melanization in peppered moth
137. A disease caused by eating fish contaminated with mercury, is called  
(a) osteosclerosis (b) minimata disease  
(c) bright's disease  
(d) hashimoto's disease
138. Although much  $\text{CO}_2$  is carried in blood, yet blood does not become acidic, because  
(a) buffer system of blood plays an important role  
(b)  $\text{CO}_2$  continuously diffuses through the tissues  
(c)  $\text{CO}_2$  combines with water to form  $\text{H}_2\text{CO}_3$ , which is neutralized by  $\text{NaCO}_3$   
(d) all of the above
139. The concept that 'population tends to increase geometrically while food supply increases arithmetically' was put forward by  
(a) Thomas Malthus (b) Adam Smith  
(c) Stuart Mill (d) Charles Darwin
140. The transgenic animals are those which have  
(a) foreign DNA in some of its cells  
(b) foreign RNA in all its cells  
(c) foreign DNA in all of its cells  
(d) both (b) and (c)
141. Which of the following metabolic disease occurs only in males ?  
(a) Lesch-Nyhan disease  
(b) Gaucher's disease  
(c) Fabry's disease  
(d) Hunter's disease
142. Hurthle cells are present in  
(a) spleen (b) liver  
(c) thyroid gland (d) lymph
143. Phylogenetic classification is one which is based on  
(a) overall similarities  
(b) common evolutionary descent  
(c) habits  
(d) utilitarian system
144. In mitochondria, cristae act as sites for  
(a) oxidation-reduction reaction  
(b) protein synthesis  
(c) breakdown of macromolecules  
(d) phosphorylation of flavoproteins
145. A product may bind to the regulatory enzyme's active site, preventing it from binding substrate and temporarily shutting down the metabolic pathway. This is called  
(a) allosteric inhibition  
(b) competitive inhibition  
(c) negative feedback  
(d) non-competitive inhibition

146. During the formation of cell wall the secreted outermost layer of cellulose is  
 (a) primary wall (b) secondary wall  
 (c) middle lamella (d) both (b) and (c)
147. Segments of DNA which are capable of moving in and out of a chromosome are termed as  
 (a) transposons (b) recon  
 (c) muton (d) replicon
148. The transition reactions  
 (a) connect glycolysis to the Krebs cycle  
 (b) give off CO<sub>2</sub>  
 (c) utilize NAD<sup>+</sup>  
 (d) include all of the above
149. Who among the following placed gymnosperms between monocots and dicots as third taxon ?  
 (a) Englar and Prantl  
 (b) Bentham and Hooker  
 (c) Hutchinson  
 (d) all of the above
150. Plasmodesmata are formed around the membranes of  
 (a) golgi bodies (b) nucleus  
 (c) chloroplast (d) none of the above
151. Which of the following statements is/are correct?  
 (a) both plasmids and viruses can serve as vectors  
 (b) vectors carry only the foreign gene into the host cells  
 (c) plasmids can carry recombinant DNA but viruses can not  
 (d) all of the above
152. In hypogynous type of flowers all floral parts arise below the  
 (a) sepals (b) gynoceium  
 (c) petals (d) androceium
153. In *Selaginella*, heterosporous spores are  
 (a) sexual and asexual  
 (b) large and small  
 (c) haploid and diploid  
 (d) all spores are of the same size
154. Which of the following plants contains mercury in their tissues ?  
 (a) *Fucus* (b) *Laminaria*  
 (c) both (a) and (b) (d) *Fusarium*
155. Protein 'canaralin' is obtained from  
 (a) carrots (b) almonds  
 (c) jack beans (d) grapes
156. Certain pollutants remain unchanged for a long time in the environment. These are not easily degradable and are termed as  
 (a) persistent (b) non-biodegradable  
 (c) both (a) and (b) (d) biodegradable
157. When the procedure of bacterial staining is carried out, the negative bacteria stain  
 (a) purple (b) red  
 (c) green (d) both (b) and (c)
158. Diatoms are placed under  
 (a) protozoans (b) fungi  
 (c) plantae (d) protista
159. A group of isodiametric cells with intercellular spaces must be  
 (a) prosenchyma (b) collenchyma  
 (c) sclerenchyma (d) parenchyma
160. Man's utilization of starch as energy source depends on the ability to convert it completely to individual glucose units. This process is initiated by the action of enzymes  
 (a) *Amylases* (b) *Cellulases*  
 (c) *Proteases* (d) none of the above
- Instructions for Q. No. 161 to 180**  
*Directions : Each of the questions given below consists of two statements, an assertions (A) and reason (R). Select the number corresponding to the appropriate response in the answer sheet as follows.*
- (a) If both assertion and reason are true and the reason is a correct explanation of the assertion  
 (b) If both assertion and reason are true but the reason is not a correct explanation of the assertion  
 (c) If the assertion is true, but the reason is false  
 (d) If both assertion and reason are false.
161. Assertion (A) : *Gibberella fujikuroi* was first called as *Fusarium moniliforme*.  
 Reason (R) : Its sexual stage was not discovered.

162. *Assertion (A)* : Tyloses are abundant in duramen.  
*Reason (R)* : They provide rigidity and strength to heartwood.
163. *Assertion (A)* : Gram, pea and mango show epigeal germination.  
*Reason (R)* : In epigeal germination radicle grows after hypocotyl.
164. *Assertion (A)* : Antitranspirants are material applied to plants for retarding transpiration.  
*Reason (R)* : Abscisic acid and phenyl mercuric acetate are not antitranspirants.
165. *Assertion (A)* : The first activity in light reaction of photosynthesis is the photolysis of  $H_2O$ .  
*Reason (R)* : PS I is not involved in the photolysis of water.
166. *Assertion (A)* : Olecranon process is present at the distal end of Ulna.  
*Reason (R)* : It articulate with the trochlea.
167. *Assertion (A)* : Mule is an example of heterosis.  
*Reason (R)* : Heterosis is the superiority of offspring to their parents.
168. *Assertion (A)* : Cardiac output is the volume of blood pumped by left or right ventricle in one minute.  
*Reason (R)* : It is calculated by multiplying the heart rate by the stroke volume.
169. *Assertion (A)* : Tea, coffee and alcohols are diuretic.  
*Reason (R)* : They suppress ADH (vasopressin).
170. *Assertion (A)* : Nephritis is the inflammation of kidney.  
*Reason (R)* : It is caused by bacterial infection.
171. *Assertion (A)* : Enzymes are protein that catalyses biochemical reactions.  
*Reason (R)* : The enzyme itself is unchanged in the reaction to take place.
172. *Assertion (A)* : Mimicry is a device adopted by the nature to protect the individuals for their own purposes.  
*Reason (R)* : It helps the animal in self defence and survival.
173. *Assertion (A)* : The sustaining surface for the gliding in certain animals, is a fold or series of folds of the skin known as patagium.  
*Reason (R)* : The gliding flights are performed by arboreal animals.
174. *Assertion (A)* : Aldosterone is a steroid hormone and is important in the control of sodium and potassium ion concentration in mammals.  
*Reason (R)* : It upgrades sodium ion concentration in the ECF by promoting reabsorption of sodium ions from renal tubules and excretion of potassium ions in urine.
175. *Assertion (A)* : Thyroid stimulating hormone is smallest polypeptide hormone of adenohipophysis of pituitary.  
*Reason (R)* : Its role is to intensify the synthesis of hormones in adrenal cortex under a direct 'feedback' regulation.
176. *Assertion (A)* : There is a gradual decrease in the energy content at successive trophic level from producer to consumer.  
*Reason (R)* : Pyramid of energy shows energy shows energy accumulation pattern at different trophic levels.
177. *Assertion (A)* : Meselson and Stahl tested the Watson and Crick theory of DNA replication.  
*Reason (R)* : They confirmed the mechanism of DNA replication by using the isotopic and centrifugation techniques.
178. *Assertion (A)* : Desired improved variety of economically useful crops are obtained by hybridization.  
*Reason (R)* : When an ovary develops into a fruit without fertilization is called hybridization.
179. *Assertion (A)* : Chromosome appears longer during leptotene.  
*Reason (R)* : The term chromosome was coined by Waldeyer.
180. *Assertion (A)* : Chromosome number is halved during Telophase-I.  
*Reason (R)* : Chromosomes whose arms are equal, termed as submetacentric.

### GENERAL KNOWLEDGE

181. 'Dazzler' is  
 (a) virus  
 (b) mascot for Cricket World Cup 2003  
 (c) bacteria  
 (d) none of these
182. Mohemmad Ali is associated with  
 (a) boxing (b) wrestling  
 (c) cricket (d) billiards
183. Who won the 2001 Miss World contest at the sun city resort in South Africa  
 (a) Zerelda Lee (b) Abgani Darego  
 (c) Diya Mirza (d) Juilet Jane Horne
184. 'Vande Matram' was taken from  
 (a) raj tarangani  
 (b) anand math  
 (c) akbar nama  
 (d) akbar khosa
185. 'Golden girl' is the biography of  
 (a) P.T.Usha  
 (b) Vijaya Lakshmi Pandit  
 (c) Sarojini Naidu  
 (d) Indira Gandhi
186. Who is the constitution head of our country?  
 (a) the President  
 (b) the Chief Justice  
 (c) the Attorney General  
 (d) the Prime Minister
187. When is the World Habitat Day?  
 (a) october 10 (b) october 8  
 (c) november 10 (d) october 3
188. Megasthenese visited India during the reign of  
 (a) Chandragupta II  
 (b) Chandragupta Maurya  
 (c) Ashoka  
 (d) Harsha
189. After returning from South Africa, Gandhiji launched his first successful 'Satyagraha' in  
 (a) chauri-chaura (b) dandi  
 (c) bardoli (d) champaran
190. The city which bore the brunt of the recent earthquake in Gujarat on January 26, 2001 is  
 (a) Ahmedabad (b) Bhuj  
 (c) Valsad (d) Gandhi Nagar
191. The Raga which is sung early in the morning is  
 (a) Todi (b) Darbari  
 (c) Bhopali (d) Bhimpalasi
192. For reproducing sound, CD (company disc) audioplayer uses a  
 (a) quartz crystal  
 (b) titanium needle  
 (c) laser beam  
 (d) barium titanate ceramic
193. Match List-1 (Books) with List II (Authors) and select the correct answer using the codes given below the Lists :
- | List I               | List II             |
|----------------------|---------------------|
| A. My Music, My Life | 1. Laxman Garkward  |
| B. Adha Gaon         | 2. Rahi Massom Raja |
| C. Radha             | 3. Ramakanta Nath   |
| D. The Pilferer      | 4. Ravi Shankar     |
| (a) A B C D          | (b) A B C D         |
| 3 2 4 1              | 4 2 3 1             |
| (c) A B C D          | (d) A B C D         |
| 4 1 3 2              | 3 1 4 2             |
194. Hiroshima day in Japan was remembered on  
 (a) August 6 (b) August 9  
 (c) October 7 (d) August 13
195. A test tube baby means fertilisation of the ovum and development taking place in the  
 (a) test tube  
 (b) uterus  
 (c) test tube and uterus respectively  
 (d) uterus and test tube respectively
196. The Bus started between India and Bangladesh flies from  
 (a) Delhi (b) Guwahati  
 (c) Calcutta (d) Asansol

197. Harappas used which type of ancient script?

- (a) symbolic
- (b) hieroglyphic
- (c) linear
- (d) pictographic

198. The fourth Buddhist council was held during the reign of

- (a) Ashoka
- (b) Chandragupta
- (c) Kanishka
- (d) Chandragupta Vikramaditya

199. Which of the following is not a computer language

- (a) BASIC
- (b) JAVA
- (c) SUMATRA
- (d) FORTRAN

200. Who played the leading role in the founding of the Indian National Congress?

- (a) A.O.Hume
- (b) Surendranath Banerjee
- (c) Gopal Krishna Gokhale
- (d) Khan Abdul Ghafer Khan

