

**Question 13.1:**

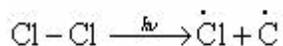
How do you account for the formation of ethane during chlorination of methane?

Answer

Chlorination of methane proceeds via a free radical chain mechanism. The whole reaction takes place in the given three steps.

**Step 1:** Initiation:

The reaction begins with the homolytic cleavage of Cl – Cl bond as:



Chlorine free radicals

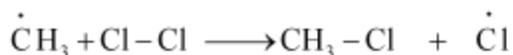
**Step 2:** Propagation:

In the second step, chlorine free radicals attack methane molecules and break down the C–H bond to generate methyl radicals as:



Methane

These methyl radicals react with other chlorine free radicals to form methyl chloride along with the liberation of a chlorine free radical.

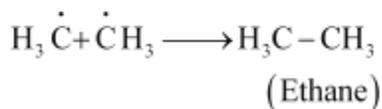


Methyl chloride

Hence, methyl free radicals and chlorine free radicals set up a chain reaction. While HCl and CH<sub>3</sub>Cl are the major products formed, other higher halogenated compounds are also formed as:

**Step 3:** Termination:

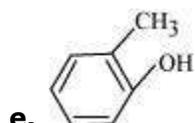
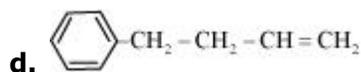
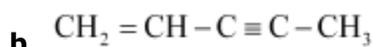
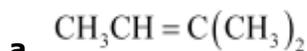
Formation of ethane is a result of the termination of chain reactions taking place as a result of the consumption of reactants as:



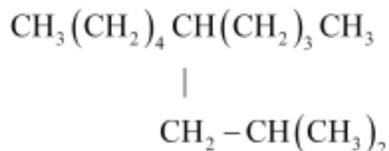
Hence, by this process, ethane is obtained as a by-product of chlorination of methane.

### Question 13.2:

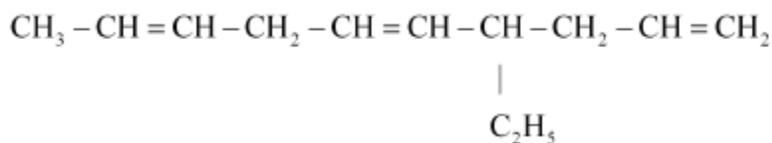
Write IUPAC names of the following compounds:



f.

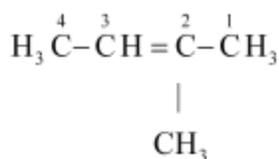


g.

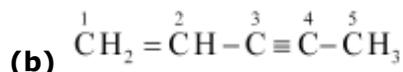


Answer

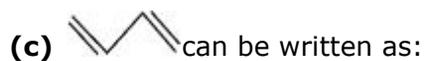
(a)



IUPAC name: 2-Methylbut-2-ene

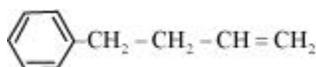


IUPAC name: Pen-1-ene-3-yne

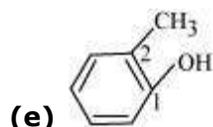


IUPAC name: 1, 3-Butadiene or Buta-1,3-diene

(d)



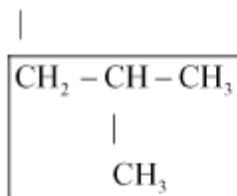
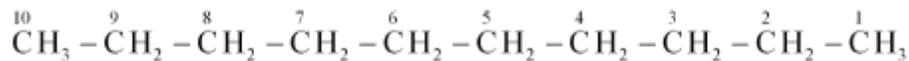
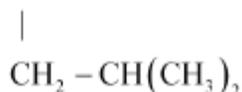
IUPAC name: 4-Phenyl but-1-ene



IUPAC name: 2-Methyl phenol

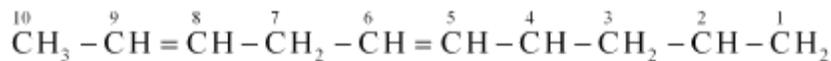
(f)

$\text{CH}_3(\text{CH}_2)_4\text{CH}(\text{CH}_2)_3\text{CH}_3$  can be written as:



IUPAC name: 5-(2-Methylpropyl)-decane

(g)



IUPAC name: 4-Ethyldeca-1, 5, 8-triene

**Question 13.3:**

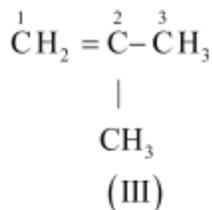
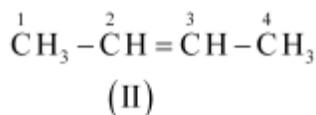
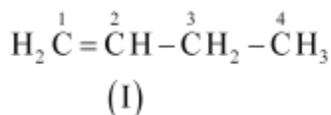
For the following compounds, write structural formulas and IUPAC names for all possible isomers having the number of double or triple bond as indicated:

**(a)** C<sub>4</sub>H<sub>8</sub> (one double bond)

**(b)** C<sub>5</sub>H<sub>8</sub> (one triple bond)

Answer

**(a)** The following structural isomers are possible for C<sub>4</sub>H<sub>8</sub> with one double bond:



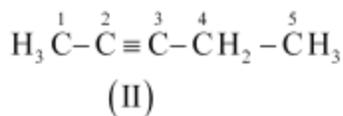
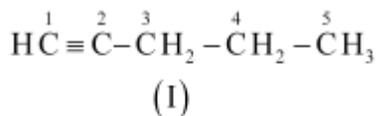
The IUPAC name of

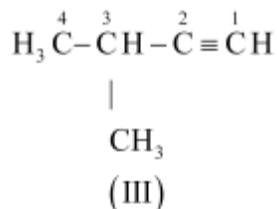
Compound (I) is But-1-ene,

Compound (II) is But-2-ene, and

Compound (III) is 2-Methylprop-1-ene.

**(b)** The following structural isomers are possible for C<sub>5</sub>H<sub>8</sub> with one triple bond:





The IUPAC name of

Compound (I) is Pent-1-yne,

Compound (II) is Pent-2-yne, and

Compound (III) is 3-Methylbut-1-ene.

#### Question 13.4:

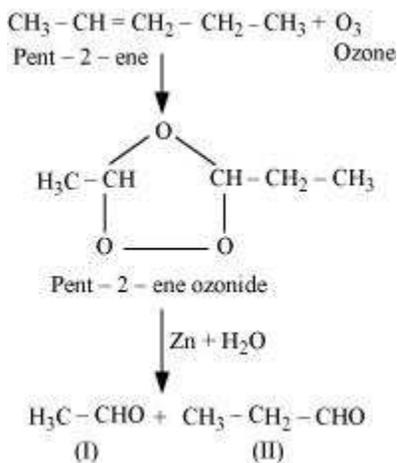
Write IUPAC names of the products obtained by the ozonolysis of the following compounds:

(i) Pent-2-ene (ii) 3,4-Dimethyl-hept-3-ene

(iii) 2-Ethylbut-1-ene (iv) 1-Phenylbut-1-ene

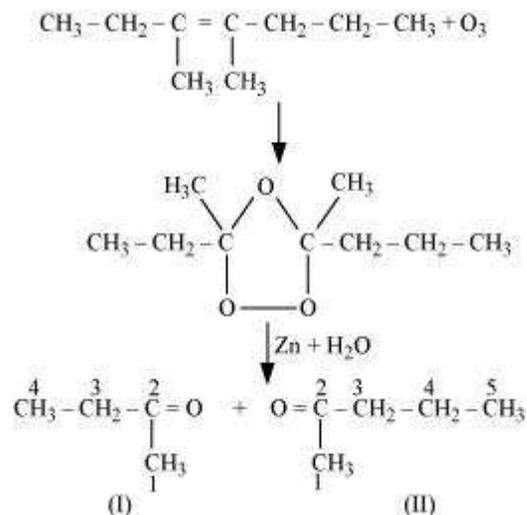
Answer

**(i)** Pent-2-ene undergoes ozonolysis as:



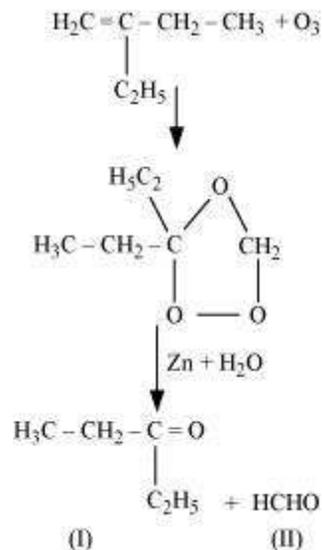
The IUPAC name of Product (I) is ethanal and Product (II) is propanal.

**(ii)** 3, 4-Dimethylhept-3-ene undergoes ozonolysis as:



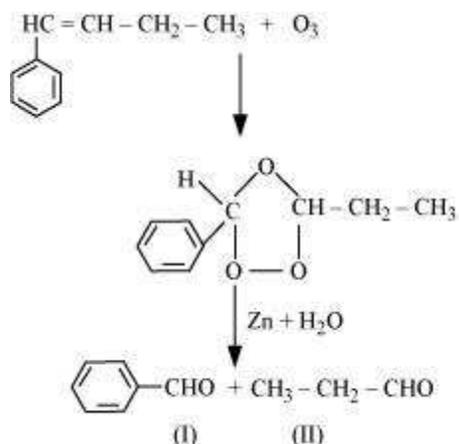
The IUPAC name of Product (I) is butan-2-one and Product (II) is Pentan-2-one.

**(iii)** 2-Ethylbut-1-ene undergoes ozonolysis as:



The IUPAC name of Product (I) is pentan-3-one and Product (II) is methanal.

**(iv)** 1-Phenylbut-1-ene undergoes ozonolysis as:

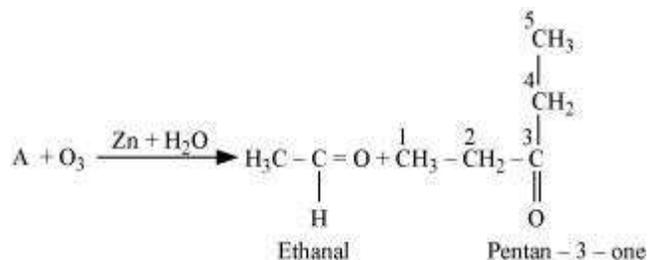


The IUPAC name of Product (I) is benzaldehyde and Product (II) is propanal.

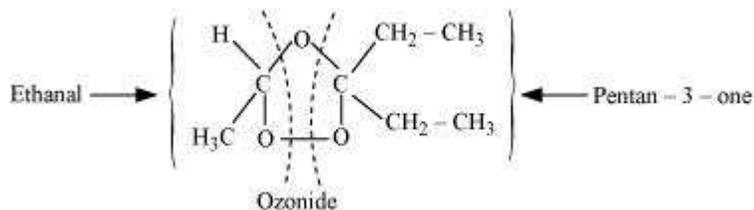
### Question 13.5:

An alkene 'A' on ozonolysis gives a mixture of ethanal and pentan-3-one. Write structure and IUPAC name of 'A'.

Answer



During ozonolysis, an ozonide having a cyclic structure is formed as an intermediate which undergoes cleavage to give the final products. Ethanal and pentan-3-one are obtained from the intermediate ozonide. Hence, the expected structure of the ozonide is:



This ozonide is formed as an addition of ozone to 'A'. The desired structure of 'A' can be obtained by the removal of ozone from the ozonide. Hence, the structural formula of 'A' is:



The final product is ethanal with molecular mass

$$= [(2 \times 12) + (4 \times 1) + (1 \times 16)]$$

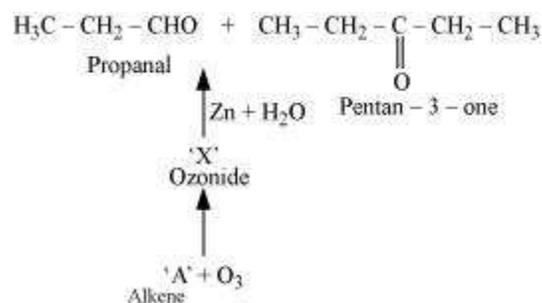
$$= 44 \text{ u}$$

**Question 13.7:**

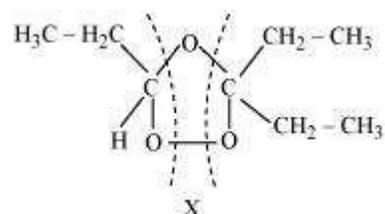
Propanal and pentan-3-one are the ozonolysis products of an alkene? What is the structural formula of the alkene?

Answer

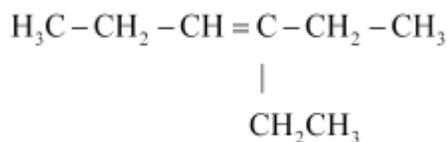
As per the given information, propanal and pentan-3-one are the ozonolysis products of an alkene. Let the given alkene be 'A'. Writing the reverse of the ozonolysis reaction, we get:



The products are obtained on the cleavage of ozonide 'X'. Hence, 'X' contains both products in the cyclic form. The possible structure of ozonide can be represented as:



Now, 'X' is an addition product of alkene 'A' with ozone. Therefore, the possible structure of alkene 'A' is:



**Question 13.8:**

Write chemical equations for combustion reaction of the following hydrocarbons:

(i) Butane (ii) Pentene (iii) Hexyne (iv) Toluene

Answer

Combustion can be defined as a reaction of a compound with oxygen.



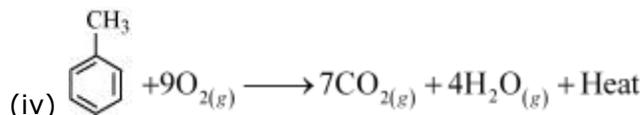
Butane



Pentene



Hexyne



Toluene

### Question 13.9:

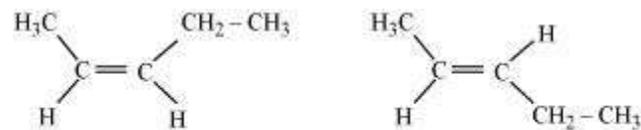
Draw the *cis* and *trans* structures of hex-2-ene. Which isomer will have higher b.p. and why?

Answer

Hex-2-ene is represented as:



Geometrical isomers of hex-2-ene are:



Cis - isomer

Trans - isomer

The dipole moment of *cis*-compound is a sum of the dipole moments of C-CH<sub>3</sub> and C-CH<sub>2</sub>CH<sub>3</sub> bonds acting in the same direction.

The dipole moment of *trans*-compound is the resultant of the dipole moments of C-CH<sub>3</sub> and C-CH<sub>2</sub>CH<sub>3</sub> bonds acting in opposite directions.

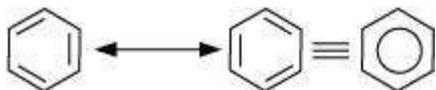
Hence, cis-isomer is more polar than trans-isomer. The higher the polarity, the greater is the intermolecular dipole-dipole interaction and the higher will be the boiling point. Hence, cis-isomer will have a higher boiling point than trans-isomer.

**Question 13.10:**

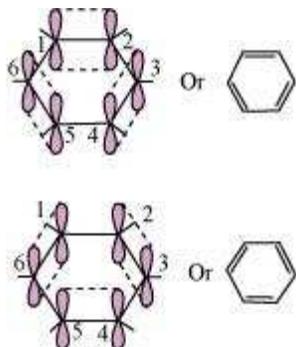
Why is benzene extra ordinarily stable though it contains three double bonds?

Answer

Benzene is a hybrid of resonating structures given as:



All six carbon atoms in benzene are  $sp^2$  hybridized. The two  $sp^2$  hybrid orbitals of each carbon atom overlap with the  $sp^2$  hybrid orbitals of adjacent carbon atoms to form six sigma bonds in the hexagonal plane. The remaining  $sp^2$  hybrid orbital on each carbon atom overlaps with the  $s$ -orbital of hydrogen to form six sigma C–H bonds. The remaining unhybridized  $p$ -orbital of carbon atoms has the possibility of forming three  $\pi$  bonds by the lateral overlap of  $C_1 - C_2, C_3 - C_4, C_5 - C_6$ , or  $C_2 - C_3, C_4 - C_5, C_6 - C_1$ .



The six  $\pi$ 's are delocalized and can move freely about the six carbon nuclei. Even after the presence of three double bonds, these delocalized  $\pi$ -electrons stabilize benzene.

**Question 13.11:**

What are the necessary conditions for any system to be aromatic?

Answer

A compound is said to be aromatic if it satisfies the following three conditions:

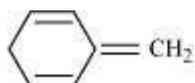
- (i) It should have a planar structure.
- (ii) The  $\pi$ -electrons of the compound are completely delocalized in the ring.

**(iii)** The total number of  $\pi$ -electrons present in the ring should be equal to  $(4n + 2)$ , where  $n = 0, 1, 2 \dots$  etc. This is known as Huckel's rule.

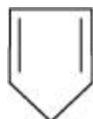
**Question 13.12:**

Explain why the following systems are **not** aromatic?

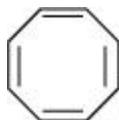
(i)



(ii)

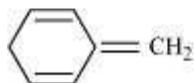


(iii)



Answer

**(i)**



For the given compound, the number of  $\pi$ -electrons is 6.

By Huckel's rule,

$$4n + 2 = 6$$

$$4n = 4$$

$$n = 1$$

For a compound to be aromatic, the value of  $n$  must be an integer ( $n = 0, 1, 2\dots$ ). Since the value of  $n$  is an integer, the given compound is aromatic in nature.

**(ii)**



For the given compound, the number of  $\pi$ -electrons is 4.

By Huckel's rule,

$$4n + 2 = 4$$

$$4n = 2$$

$$n = \frac{1}{2}$$

For a compound to be aromatic, the value of  $n$  must be an integer ( $n = 0, 1, 2, \dots$ ), which is not true for the given compound. Hence, it is not aromatic in nature.

**(iii)**



For the given compound, the number of  $\pi$ -electrons is 8.

By Huckel's rule,

$$4n + 2 = 8$$

$$4n = 6$$

$$n = \frac{3}{2}$$

For a compound to be aromatic, the value of  $n$  must be an integer ( $n = 0, 1, 2, \dots$ ). Since the value of  $n$  is not an integer, the given compound is not aromatic in nature.

**Question 13.13:**

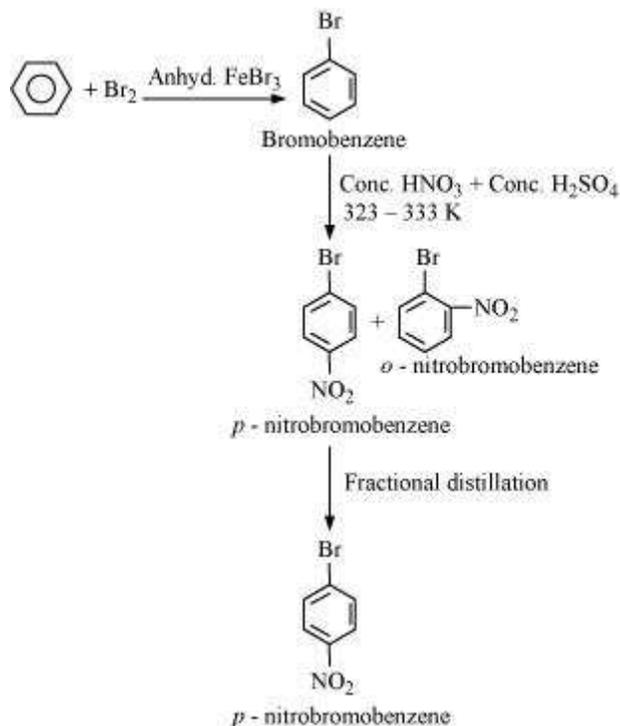
How will you convert benzene into

(i) *p*-nitrobromobenzene (ii) *m*-nitrochlorobenzene

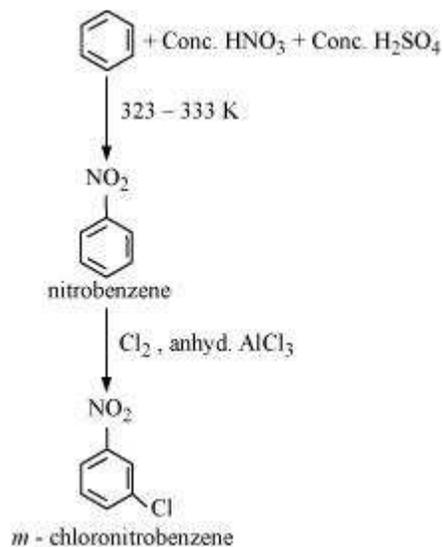
(iii) *p*-nitrotoluene (iv) acetophenone

Answer

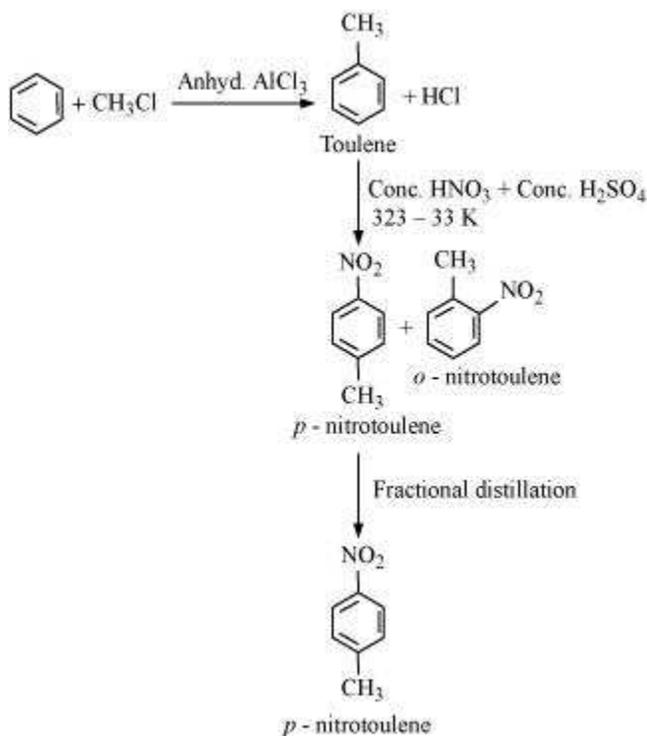
**(i)** Benzene can be converted into *p*-nitrobromobenzene as:



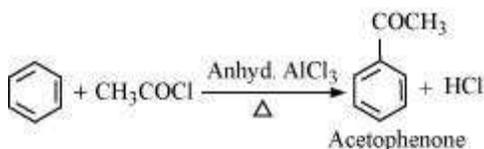
(ii) Benzene can be converted into *m*-nitrochlorobenzene as:



(iii) Benzene can be converted into *p*-nitrotoluene as:



(iv) Benzene can be converted into acetophenone as:



#### Question 13.14:

In the alkane  $\text{H}_3\text{C}-\text{CH}_2-\text{C}(\text{CH}_3)_2-\text{CH}_2-\text{CH}(\text{CH}_3)_2$ , identify  $1^\circ, 2^\circ, 3^\circ$  carbon atoms and give the number of H atoms bonded to each one of these.

Answer



$1^\circ$  carbon atoms are those which are bonded to only one carbon atom i.e., they have only one carbon atom as their neighbour. The given structure has five  $1^\circ$  carbon atoms and fifteen hydrogen atoms attached to it.

2° carbon atoms are those which are bonded to two carbon atoms i.e., they have two carbon atoms as their neighbours. The given structure has two 2° carbon atoms and four hydrogen atoms attached to it.

3° carbon atoms are those which are bonded to three carbon atoms i.e., they have three carbon atoms as their neighbours. The given structure has one 3° carbon atom and only one hydrogen atom is attached to it.

### Question 13.15:

What effect does branching of an alkane chain has on its boiling point?

Answer

Alkanes experience inter-molecular Van der Waals forces. The stronger the force, the greater will be the boiling point of the alkane.

As branching increases, the surface area of the molecule decreases which results in a small area of contact. As a result, the Van der Waals force also decreases which can be overcome at a relatively lower temperature. Hence, the boiling point of an alkane chain decreases with an increase in branching.

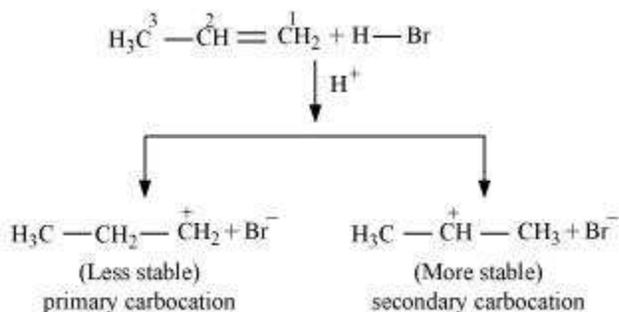
### Question 13.16:

Addition of HBr to propene yields 2-bromopropane, while in the presence of benzoyl peroxide, the same reaction yields 1-bromopropane. Explain and give mechanism.

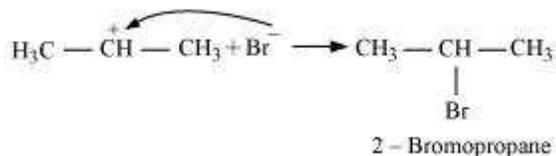
Answer

Addition of HBr to propene is an example of an electrophilic substitution reaction.

Hydrogen bromide provides an electrophile, H<sup>+</sup>. This electrophile attacks the double bond to form 1° and 2° carbocations as shown:

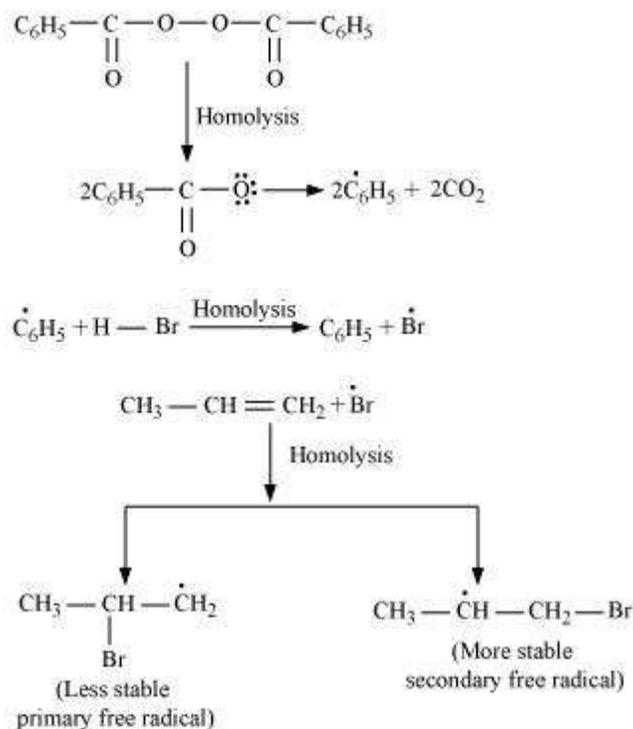


Secondary carbocations are more stable than primary carbocations. Hence, the former predominates since it will form at a faster rate. Thus, in the next step,  $\text{Br}^-$  attacks the carbocation to form 2 - bromopropane as the major product.

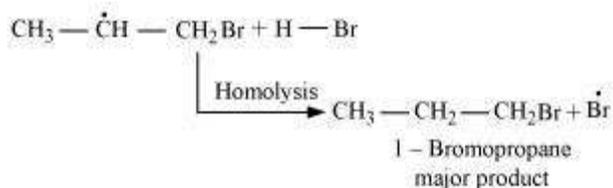


This reaction follows Markovnikov's rule where the negative part of the addendum is attached to the carbon atom having a lesser number of hydrogen atoms.

In the presence of benzoyl peroxide, an addition reaction takes place anti to Markovnikov's rule. The reaction follows a free radical chain mechanism as:



Secondary free radicals are more stable than primary radicals. Hence, the former predominates since it forms at a faster rate. Thus, 1 - bromopropane is obtained as the major product.



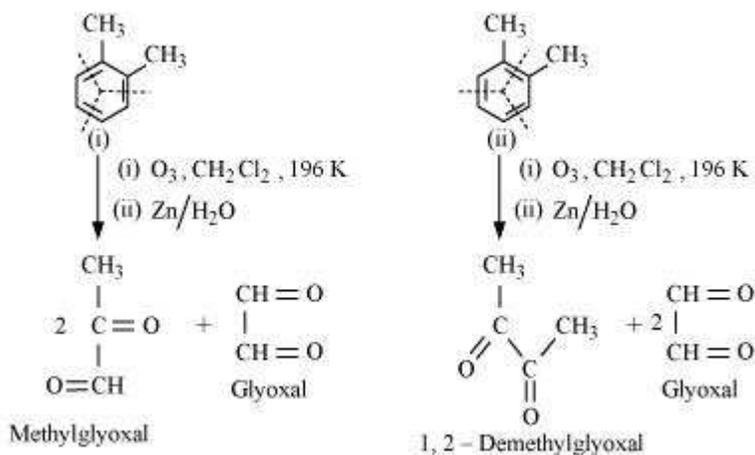
In the presence of peroxide, Br free radical acts as an electrophile. Hence, two different products are obtained on addition of HBr to propene in the absence and presence of peroxide.

### Question 13.17:

Write down the products of ozonolysis of 1, 2-dimethylbenzene (*o*-xylene). How does the result support Kekulé structure for benzene?

Answer

*o*-xylene has two resonance structures:



All three products, i.e., methyl glyoxal, 1, 2-demethylglyoxal, and glyoxal are obtained from two Kekule structures. Since all three products cannot be obtained from any one of the two structures, this proves that *o*-xylene is a resonance hybrid of two Kekule structures (I and II).

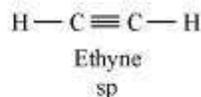
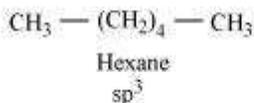
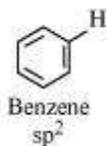
### Question 13.18:

Arrange benzene, *n*-hexane and ethyne in decreasing order of acidic behaviour. Also give reason for this behaviour.

Answer

Acidic character of a species is defined on the basis of ease with which it can lose its H-atoms.

The hybridization state of carbon in the given compound is:



As the  $s$ -character increases, the electronegativity of carbon increases and the electrons of C-H bond pair lie closer to the carbon atom. As a result, partial positive charge of H-atom increases and  $\text{H}^+$  ions are set free.

The  $s$ -character increases in the order:



Hence, the decreasing order of acidic behaviour is Ethyne > Benzene > Hexane.

### Question 13.19:

Why does benzene undergo electrophilic substitution reactions easily and nucleophilic substitutions with difficulty?

Answer

Benzene is a planar molecule having delocalized electrons above and below the plane of ring. Hence, it is electron-rich. As a result, it is highly attractive to electron deficient species i.e., electrophiles.

Therefore, it undergoes electrophilic substitution reactions very easily. Nucleophiles are electron-rich. Hence, they are repelled by benzene. Hence, benzene undergoes nucleophilic substitutions with difficulty.

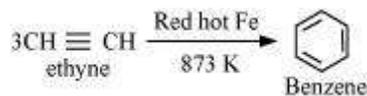
### Question 13.20:

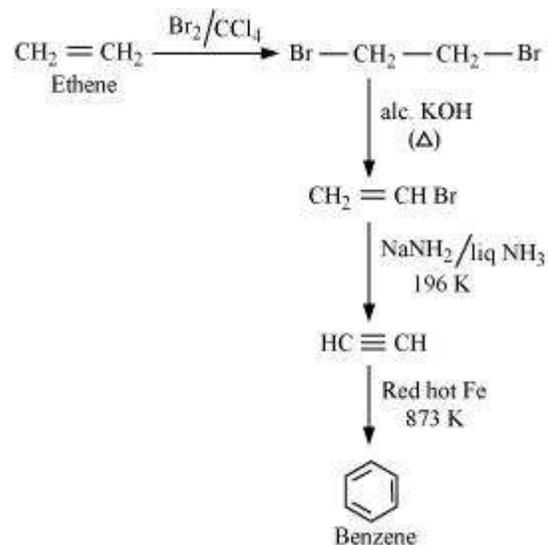
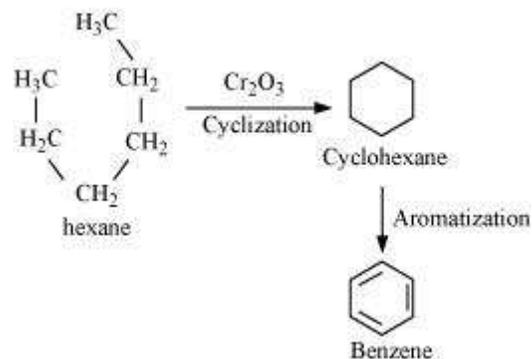
How would you convert the following compounds into benzene?

(i) Ethyne (ii) Ethene (iii) Hexane

Answer

(i) Benzene from Ethyne:

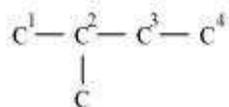


**(ii) Benzene from Ethene:****(iii) Hexane to Benzene****Question 13.21:**

Write structures of all the alkenes which on hydrogenation give 2-methylbutane.

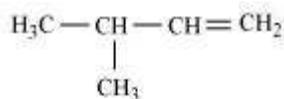
Answer

The basic skeleton of 2-methylbutane is shown below:

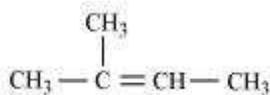


On the basis of this structure, various alkenes that will give 2-methylbutane on hydrogenation are:

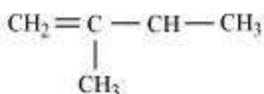
(a)



(b)



(c)

**Question 13.22:**

Arrange the following set of compounds in order of their decreasing relative reactivity with an electrophile,  $\text{E}^+$

(a) Chlorobenzene, 2,4-dinitrochlorobenzene, *p*-nitrochlorobenzene

(b) Toluene, *p*- $\text{H}_3\text{C}-\text{C}_6\text{H}_4-\text{NO}_2$ , *p*- $\text{O}_2\text{N}-\text{C}_6\text{H}_4-\text{NO}_2$ .

Answer

Electrophiles are reagents that participate in a reaction by accepting an electron pair in order to bond to nucleophiles.

The higher the electron density on a benzene ring, the more reactive is the compound towards an electrophile,  $\text{E}^+$  (Electrophilic reaction).

(a) The presence of an electron withdrawing group (i.e.,  $\text{NO}_2^-$  and  $\text{Cl}^-$ ) deactivates the aromatic ring by decreasing the electron density.

Since  $\text{NO}_2^-$  group is more electron withdrawing (due to resonance effect) than the  $\text{Cl}^-$  group (due to inductive effect), the decreasing order of reactivity is as follows:

Chlorobenzene > *p*-nitrochlorobenzene > 2, 4 - dinitrochlorobenzene

(b) While  $\text{CH}_3^-$  is an electron donating group,  $\text{NO}_2^-$  group is electron withdrawing. Hence, toluene will have the maximum electron density and is most easily attacked by  $\text{E}^+$ .

$\text{NO}_2^-$  is an electron withdrawing group. Hence, when the number of  $\text{NO}_2^-$  substituents is greater, the order is as follows:

Toluene > *p*- $\text{CH}_3-\text{C}_6\text{H}_4-\text{NO}_2$ , *p*- $\text{O}_2\text{N}-\text{C}_6\text{H}_4-\text{NO}_2$

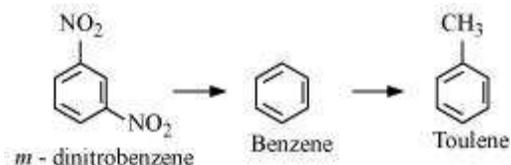
**Question 13.23:**

Out of benzene, *m*-dinitrobenzene and toluene which will undergo nitration most easily and why?

Answer

The ease of nitration depends on the presence of electron density on the compound to form nitrates. Nitration reactions are examples of electrophilic substitution reactions where an electron-rich species is attacked by a nitronium ion ( $\text{NO}_2^+$ ).

Now,  $\text{CH}_3$ - group is electron donating and  $\text{NO}_2$ - is electron withdrawing. Therefore, toluene will have the maximum electron density among the three compounds followed by benzene. On the other hand, *m*-Dinitrobenzene will have the least electron density. Hence, it will undergo nitration with difficulty. Hence, the increasing order of nitration is as follows:

**Question 13.24:**

Suggest the name of a Lewis acid other than anhydrous aluminium chloride which can be used during ethylation of benzene.

Answer

The ethylation reaction of benzene involves the addition of an ethyl group on the benzene ring. Such a reaction is called a Friedel-Craft alkylation reaction. This reaction takes place in the presence of a Lewis acid.

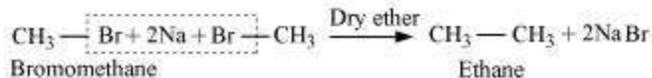
Any Lewis acid like anhydrous  $\text{FeCl}_3$ ,  $\text{SnCl}_4$ ,  $\text{BF}_3$  etc. can be used during the ethylation of benzene.

**Question 13.25:**

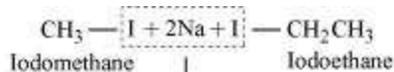
Why is Wurtz reaction not preferred for the preparation of alkanes containing odd number of carbon atoms? Illustrate your answer by taking one example.

Answer

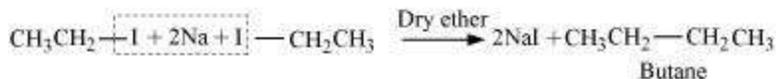
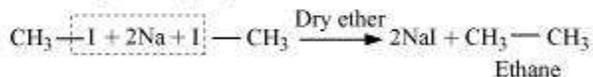
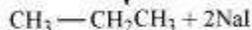
Wurtz reaction is limited for the synthesis of symmetrical alkanes (alkanes with an even number of carbon atoms) In the reaction, two similar alkyl halides are taken as reactants and an alkane, containing double the number of carbon atoms, are formed. Example:



Wurtz reaction cannot be used for the preparation of unsymmetrical alkanes because if two dissimilar alkyl halides are taken as the reactants, then a mixture of alkanes is obtained as the products. Since the reaction involves free radical species, a side reaction also occurs to produce an alkene. For example, the reaction of bromomethane and iodoethane gives a mixture of alkanes.



↓  
Dry ether



The boiling points of alkanes (obtained in the mixture) are very close. Hence, it becomes difficult to separate them