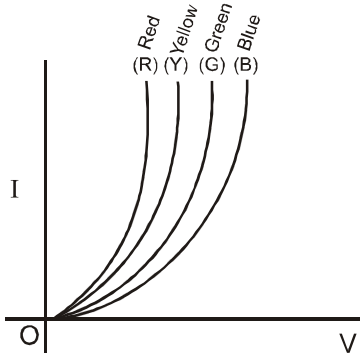
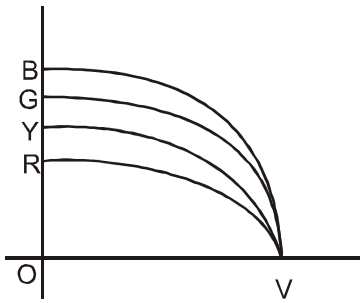
24. The I – V characteristic of an LED is :

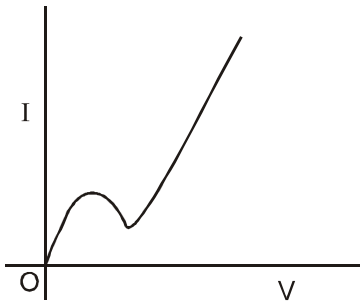
(1)



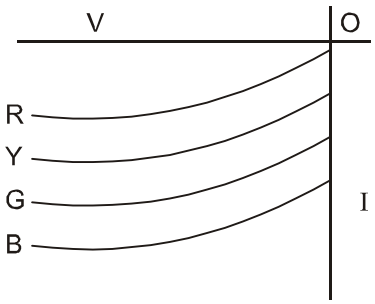
(2)



(3)



(4)



Sol. For same value of current higher value of voltage is required for higher frequency hance (1) should be correct answers.

Ans (1)

25. Assume that a drop of liquid evaporates by decrease in its surface energy, so that its temperature remains unchanged. What should be the minimum radius of the drop for this to be possible ? The surface tension is T, density of liquid is ρ and L is its latent heat of vaporization.

(1) $\rho L/T$ (2) $\sqrt{T/\rho L}$ (3) $T/\rho L$ (4) $2T/\rho L$

Sol. When radius is decrease by dr

decrease in surface energy = Heat required for vaporisation

$(4\pi r dr) \times T \times 2 = 4\pi r^2 dr\, \rho \qquad \Rightarrow r = \dfrac{2T}{\rho L}$

Ans. (4)

26. Ina hydrogen like atom electron make transition from an energy level with quantum number n to another with quantum number (n–1). If n>>1, the frequency of radiation emitted is proportional to :

(1) $\dfrac{1}{n}$ (2) $\dfrac{1}{n^2}$ (3) $\dfrac{1}{n^{3/2}}$ (4) $\dfrac{1}{n^3}$

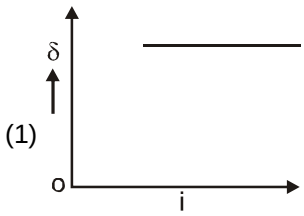
Sol. $\Delta E = h\nu$

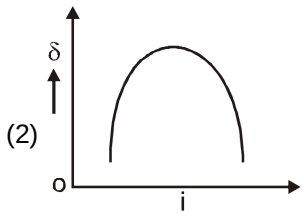
$$\nu = \frac{\Delta E}{h} = k\left[\frac{1}{(n-1)^2} - \frac{1}{n^2}\right] = \frac{k2n}{n^2(n-1)^2}$$

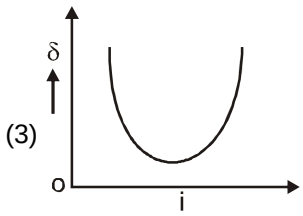
$$\approx \frac{2k}{n^3} \propto \frac{1}{n^3}$$

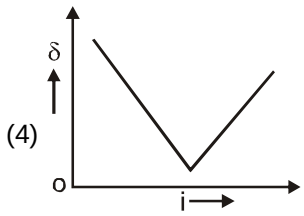
Ans. (4)

27. The graph between angle of deviation (δ) and angle of incidence (i) for a triangular prism is represented by :

(1) 

(2) 

(3) 

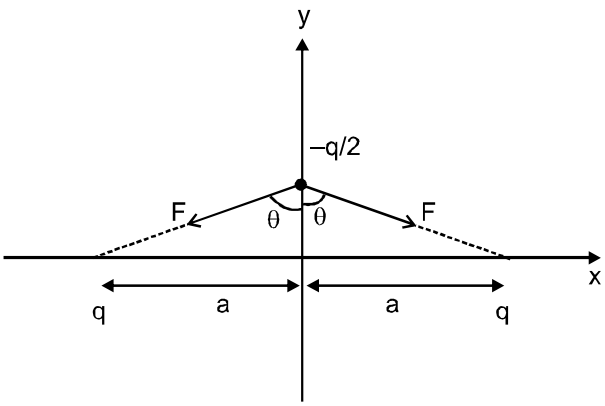
(4) 

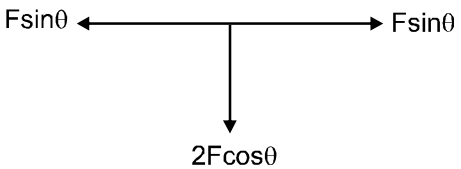
Sol. **Ans. (3)**

28. Two charges, each equal to q, are kept at x = – a and x = a on the x-axis. A particle of mass m and charge $q_0 = \frac{q}{2}$ is placed at the origin. If charge q_0 is given a small displacement (y <<a) along the y-axis, the net force acting on the particle is proportional to :

(1) y (2) –y (3) $\frac{1}{y}$ (4) $-\frac{1}{y}$

Sol.



$\Rightarrow$ 

$\Rightarrow F_{net} = 2F\cos\theta$

$$F_{net} = \frac{2kq\left(\frac{q}{2}\right)}{\left(\sqrt{y^2+a^2}\right)^2} \cdot \frac{y}{\sqrt{y^2+a^2}}$$

$$F_{net} = \frac{2kq\left(\frac{q}{2}\right)y}{(y^2+a^2)^{3/2}} \quad \Rightarrow \quad \frac{kq^2y}{a^3} \propto y$$

Ans. (1)

29. Two short bar magnets of length 1 cm each have magnetic moments 1.20 Am^2 and 1.00 Am^2 respectively. They are placed on a horizontal table parallel to each other with their N poles poining towards the South. They have a common magnetic equator and are separated by a distance of 20.0 cm. The value of the resultand horizontal magnetic induction at the mid - point O of the line joining their centres is close to

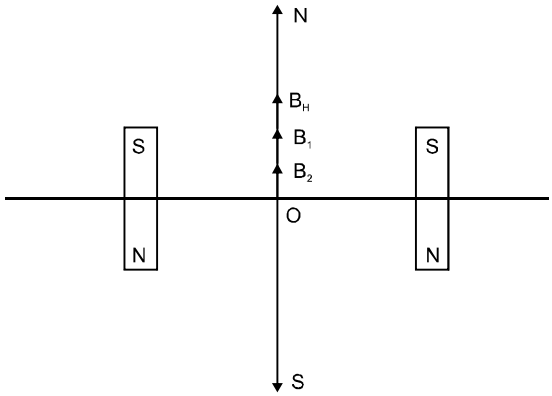
(Horizontal component of earth's magnetic induction is 3.6×10^{-5} Wb/m^2)

(1) 3.6×10^{-5} Wb/m^2 (2) 2.56×10^{-4} Wb/m^2 (3) 3.50×10^{-4} Wb/m^2 (4) 5.80×10^{-4} Wb/m^2

Sol. $B_{net} = B_1 + B_2 + B_H$

$$B_{net} = \frac{\mu_0}{4\pi}\frac{(M_1+M_2)}{r^3} + B_H$$

$$= \frac{10^{-7}(1.2+1)}{(0.1)^3} + 3.6 \times 10^{-5} = 2.56 \times 10^{-4} \text{ wb/m}^2$$

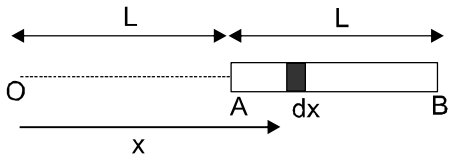


Ans. (2)

30. A charge Q is uniformly distributed over a long rod AB of length L as shown in the figure. The electric potential at the point O lying at distance L from the end A is :

(1) $\frac{Q}{8\pi \in_0 L}$ (2) $\frac{3Q}{4\pi \in_0 L}$ (3) $\frac{Q}{4\pi \in_0 L \ln 2}$ (4) $\frac{Q \ln 2}{4\pi \in_0 L}$

Sol.



$$V = \int_L^{2L} \frac{kdq}{x}$$

$$= \int_L^{2L} \frac{1}{4\pi\varepsilon_0}\frac{\left(\frac{q}{L}\right)dx}{x}$$

$$= \frac{q}{4\pi\varepsilon_0 L}\ell n(2)$$

Ans. (4)

PART B – CHEMISTRY

31. Which of the following complex species is not expected to exhibit optical isomerism ?

(1) $[Co(en)_3]^{3+}$ (2) $[Co(en)_2 Cl_2]^+$

(3) $[Co(NH_3)_3 Cl_3]$ (4) $[Co(en)(NH_3)_2 Cl_2]^+$

Ans. (3)

Sol. $[Co(NH_3)_3Cl_3]$ show facial as well as meridional isomerism. But both contain plane of symmetry.

So, the answer is (3).

32. Which one of the following molecules is expected to exhibit diamagnetic behaviour ?

(1) C_2 (2) N_2 (3) O_2 (4) S_2

Ans. **(1) and (2)**

Sol. N_2 and C_2 both are diamagnetic

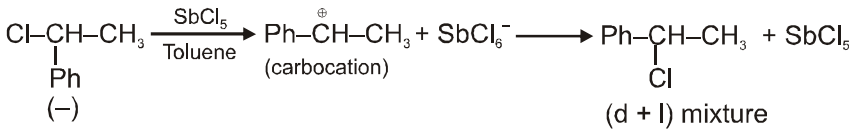
Ans is (1) and (2).

33. A solution of (–) – 1 – chloro–1–phenylethane in toluene racemises slowly in the presence of a small amount of $SbCl_5$, due to the formation of :

(1) carbanion (2) carbene

(3) carbocation (4) free radical

Ans. (3)

Sol. 

34. Given : $E^0_{Cr^{3+}/Cr} = -0.74$ V ; $E^0_{MnO_4^-/Mn^{2+}} = 1.51$ V

$E^0_{Cr_2O_7^{2-}/Cr^{3+}} = 1.33$ V ; $E^0_{Cl/Cl^-} = 1.36$ V

Based on the data given above, strongest oxidising agent will be :

(1) Cl (2) Cr^{3+} (3) Mn^{2+} (4) MnO_4^-

Ans. (4)

Sol. Higher the SRP, better is oxidising agent

Hence MnO_4^- is strongest oxidising agent.

35. A piston filled with 0.04 mol of an ideal gas expands reversibly from 50.0 mL to 375 mL at a constant temperature of 37.0ºC. As it does so, it absorbs 208 J of heat. The values of q and w for the process will be: (R = 8.314 J/mol K) (ln 7.5 = 2.01)

(1) q = +208 J, w = – 208 J (2) q = –208 J, w = – 208 J

(3) q = –208 J, w = + 208 J (4) q = +208 J, w = + 208 J

Ans. (1)

Sol. The process is isothermal expansion Hence, q = - w $\Delta u = 0$

q = + 208 J

w = – 208 J(expansion work)

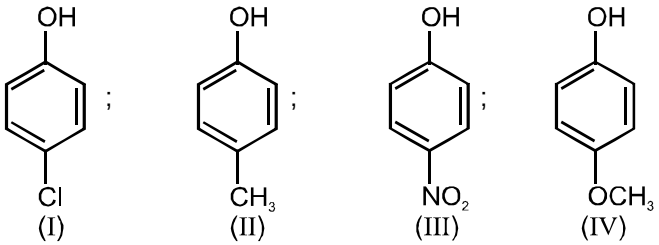
36. The molarity of a solution obtained by mixing 750 mL of 0.5(M) HCl with 250 mL of 2(M)HCl will be :

(1) 0.875 M (2) 1.00 M (3) 1.75 M (4) 0.975 M

Ans. (1)

Sol. $M_f = \frac{M_1V_1 + M_2V_2}{V_1 + V_2} = \frac{0.5 \times \frac{3}{4} + 2 \times \frac{1}{4}}{1}$ = 0.875 M.

37. Arrange the following compounds in order of decreasing acidity :

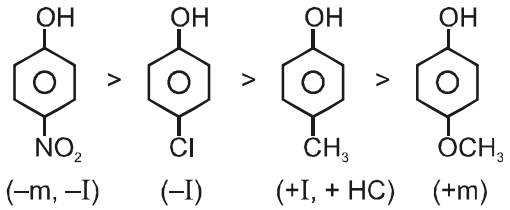


(1) II > IV > I > III (2) I > II > III > IV

(3) III > I > II > IV (4) IV > III > I > II

Ans. (3)

Sol.



electron releasing group decreases and electron withdrawing group increases acidic strength.

38. For gaseous state, if most probable speed is denoted by C*, average speed by $\overline{C}$ and mean square speed by C, then for a large number of molecules the ratios of these speeds are :

(1) C* : $\overline{C}$: C = 1.225 : 1.128 : 1 (2) C* : $\overline{C}$: C = 1.128 : 1.225 : 1

(3) C* : $\overline{C}$: C = 1 : 1.128 : 1.225 (4) C* : $\overline{C}$: C = 1 : 1.225 : 1.128

Ans. (3)

Sol. C* = most probable speed = $\sqrt{\frac{2RT}{M}}$

$\overline{C}$ = average speed = $\sqrt{\frac{8RT}{\pi M}}$

C = Mean square speed corrected as rms = $\sqrt{\frac{3RT}{M}}$

$C^* < \overline{C} < C$

$C^* : \overline{C} : C = 1 : \sqrt{\frac{4}{\pi}} : \sqrt{\frac{3}{2}}$ = 1 : 1.128 : 1.225

Note : As no option correspond to mean square speed, it is understood as misprint. It should be root means square speed

So, **Ans is (3)**

39. The rate of a reaction doubles when its temperature changes from 300 K to 310 K. Activation energy of such a reaction will be : (R = 8.314 JK^{-1} mol^{-1} and log 2 = 0.301)

(1) 53.6 kJ mol^{-1} (2) 48.6 kJ mol^{-1}

(3) 58.5 kJ mol^{-1} (4) 60.5 kJ mol^{-1}

Ans. (1)

Sol. $$\log\frac{K_2}{K_1} = \frac{-E_a}{2.030R}\left(\frac{1}{T_2} - \frac{1}{T_1}\right)$$

$$\frac{K_2}{K_1} = 2 \;\; ; \;\; T_2 = 310 \text{ K} \qquad T_1 = 300 \text{ K}$$

$$\Rightarrow \log 2 = \frac{-E_a}{2.303 \times 8.134}\left(\frac{1}{310} - \frac{1}{300}\right)$$

$$\Rightarrow E_a = 53598.6 \text{ J/mol} = 53.6 \text{ KJ/mol}$$

Ans is (1)

40. A compound with molecular mass 180 is acylated with CH_3COCl to get a compound with molecular mass 390. The number of amino groups present per molecule of the former compound is :

(1) 2 (2) 5

(3) 4 (4) 6

Ans. (2)

Sol. By reaction with one mole of $CH_3-\overset{\overset{O}{\|}}{C}-Cl$ with one $-NH_2$ group the molecular mass increases with 42 unit. Since the mass increases by (390 – 180) = 210 hence the number of $-NH_2$ groups is 5.

$$R-NH_2 + CH_3-\overset{\overset{O}{\|}}{C}-Cl \xrightarrow[(-HCl)]{} R-NH-\overset{\overset{O}{\|}}{C}-CH_3$$

41. Which of the following arrangements does not represent the correct order of the property stated against it ?

(1) $V^{2+} < Cr^{2+} < Mn^{2+} < Fe^{2+}$: paramagnetic behaviour

(2) $Ni^{2+} < Co^{2+} < Fe^{2+} < Mn^{2+}$: ionic size

(3) $Co^{3+} < Fe^{3+} < Cr^{3+} < Sc^{3+}$: stability in aqueous solution

(4) Sc < Ti < Cr < Mn : number of oxidation states

Ans. (1)

Sol. (1) V^{2+} = 3 unpaired electrons

Cr^{2+} = 4 unpaired electrons

Mn^{2+} = 5 unpaired electrons

Fe^{2+} = 4 unpaired electrons

Hence the order of paramagnetic behaviour should be

$V^{2+} < Cr^{2+} = Fe^{2+} < Mn^{2+}$

(2) ionic size decrease from left to right in same period

(3) As per data from NCERT.

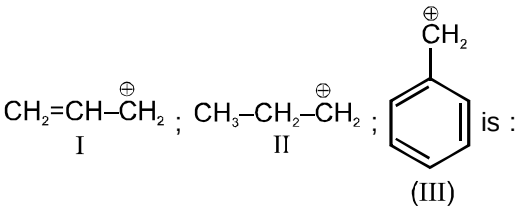
Co^{3+}/Co^{2+}= 1.97 ; Fe^{3+}/Fe^{2+} =0.77 ; Cr^{3+}/Cr^{2+}= - 0.41

Sc^{3+} is highly stable (It does not show +2)

(4) The oxidation states increases as we go from group 3 to group 7 in same period.

Ans is (1)

42. The order of stability of the following carbocations :

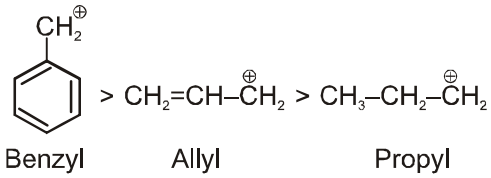


(1) III > II > I (2) II > III > I

(3) I > II > III (4) III > I > II

Ans. (4)

Sol. The order of stability of carbocation will be



43. Consider the following reaction :

$$xMnO_4^- + yC_2O_4^{2-} + zH^+ \rightarrow xMn^{2+} + 2yCO_2 + \frac{z}{2}H_2O$$

The values of x, y and z in the reaction are, respectively :

(1) 5, 2 and 16 (2) 2, 5 and 8

(3) 2, 5 and 16 (4) 5, 2 and 8

Ans. (3)

Sol. $MnO_4^- + C_2O_4^{2-} + H^+ \longrightarrow Mn^2 + CO_2 + H_2O$

vf = 1(7 –2) vf = 2(3–2)

= 5 = 2

$\therefore$ Balanced Equation :

$2MnO_4^- + 5C_2O_4^{2-} + 16\ H^+ \longrightarrow 2Mn^{2+} + 10\ CO_2 + 8H_2O$

So, x = 2, y = 5 & z = 16.

44. Which of the following is the wrong statement ?

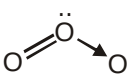
(1) ONCl and ONO^- are not isoelectronic. (2) O_3 molecule is bent

(3) Ozone is violet-black in solid state (4) Ozone is diamagnetic gas.

Ans. (All statement are correct there is no answer).

Sol. (1) $ONCl = 8 + 7 + 17 = 32e^-$

$ONO^- = 8 + 7 + 8 + 1 = 24e^-$ (correct)

(2)  Central atom O is sp^2 hybridised with 1 lone pair, so bent shape (correct)

(3) Ozone is violet-black in solid state. (Ref. NCERT & shriver atkins)

(4) O_3 has no unpaired electrons, so diamagnetic (correct)

45. A gaseous hydrocarbon gives upon combustion 0.72 g of water and 3.08 g. of CO_2. The empirical formula of the hydrocarbon is :

(1) C_2H_4 (2) C_3H_4

(3) C_6H_5 (4) C_7H_8

Ans. (4)

Sol. 18g H_2O contains 2g H

$\therefore$ 0.72 g H_2O contains 0.08 gH.

44 g CO_2 contains 12g C

$\therefore$ 3.08 g CO_2 contains 0.84 g C

$\therefore$ C : H = $\frac{0.84}{12}:\frac{0.08}{1}$ = 0.07 : 0.08

= 7 : 8

$\therefore$ Empirical formula = C_7H_8

46. In which of the following pairs of molecules/ions, both the species are not likely to exist ?

(1) H_2^+, He_2^{2-} (2) H_2^-, He_2^{2-}

(3) H_2^{2+}, He_2 (4) H_2^-, He_2^{2+}

Ans. (3)

Sol. H_2^{2+} : Bond order = 0

He_2 : Bond order = $\frac{2-2}{2}=0$

So, both H_2^{2+} & He_2 do not exist.

47. Which of the following exists as covalent crystals in the solid state ?

(1) Iodine (2) Silicon

(3) Sulphur (4) Phosphorus

Ans. (2)

Sol. Silicon exists as covalent crystal in solid state. (Network like structure, like diamond).

48. Synthesis of each molecule of glucose in photosynthesis involves :

(1) 18 molecules of ATP (2) 10 molecules of ATP

(3) 8 molecules of ATP (4) 6 molecules of ATP

Ans. (1) Fact

Sol. $6CO_2 + 12NADPH + 18\,ATP \longrightarrow C_6H_{12}O_6 + 12NADP + 18\,ADP$

49. The coagulating power of electrolytes having ions Na^+, Al^{3+} and Ba^{2+} for arsenic sulphide sol increases in the order :

(1) $Al^{3+} < Ba^{2+} < Na^+$ (2) $Na^+ < Ba^{2+} < Al^{3+}$

(3) $Ba^{2+} < Na^+ < Al^{3+}$ (4) $Al^{3+} < Na^+ < Ba^{2+}$

Ans. (2)

Sol. According to Hardy Schulze rule, greater the charge on cation, greater is its coagulating power for negatively charged solution. So, order of coagulating power : $Na^+ < Ba^{2+} < Al^{3+}$.

50. Which of the following represents the correct order of increasing first ionization enthalpy for Ca, Ba, S, Se and Ar ?

(1) Ca < S < Ba < Se < Ar (2) S < Se < Ca < Ba < Ar

(3) Ba < Ca < Se < S < Ar (4) Ca < Ba < S < Se < Ar

Ans. (3)

Sol. Order of increasing ΔH_{IE_1} : Ba < Ca < Se < S < Ar

Ba < Ca ; Se < S : On moving top to bottom in a group, size increases. So ionisation energy decreases.
Ar : Maximum value of ionisation energy, since it is an inert gas.

51. Energy of an electron is given by $E = -2.178 \times 10^{-18} J\left(\frac{Z^2}{n^2}\right)$. Wavelength of light required to excite an electron in an hydrogen atom from level n = 1 to n = 2 will be :

($h = 6.62 \times 10^{-34}$ Js and $c = 3.0 \times 10^{8}$ ms^{-1})

(1) 1.214×10^{-7} m (2) 2.816×10^{-7} m

(3) 6.500×10^{-7} m (4) 8.500×10^{-7} m

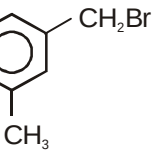
Ans. (1)

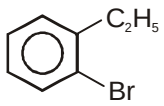
Sol. $\Delta E = 2.178 \times 10^{-18}\left(\frac{1}{1^2}-\frac{1}{2^2}\right)=\frac{hC}{\lambda}$

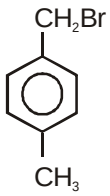
$$2.178 \times 10^{-18}\left(\frac{1}{1^2}-\frac{1}{2^2}\right)=\frac{6.62\times10^{-34}\times3.0\times10^{8}}{\lambda}$$

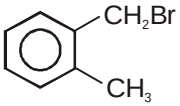
$\therefore \lambda \approx 1.214 \times 10^{-7}$m

52. Compound (A), C_8H_9Br, gives a white precipitate when warmed with alcoholic $AgNO_3$. Oxidation of (A) gives an acid (B), $C_8H_6O_4$. (B) easily forms anhydride on heating. Identify the compound (A).

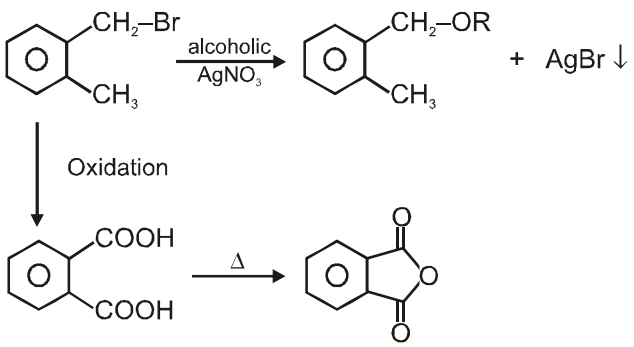
(1) 

(2) 

(3) 

(4) 

Ans. (4)

Sol. 

(Phthalic anhydride)

53. Four successive members of the first row transition elements are listed below with atomic numbers. Which one of them is expected to have the highest $E^0_{M^{3+}/M^{2+}}$ value ?

(1) Cr(Z = 24) (2) Mn(Z = 25)

(3) Fe(Z = 26) (4) Co(Z = 27)

Ans. (4)

Sol. $E^o_{Cr^{3+}/Cr^{2+}} = -0.41$ V ; $E^o_{Mn^{3+}/Mn^{2+}} = +1.57$ V ; $E^o_{Fe^{3+}/Fe^{2+}} = +0.77$ V ; $E^o_{Co^{3+}/Co^{2+}} = +1.97$ V

SRP value normaly increases from left to right in the period of d-block elements. Some SRP value are exceptionally higher due to stability of product ion. For e.g. $E^o_{Mn^{3+}/Mn^{2+}} = +1.57$ V ; $E^o_{Co^{3+}/Co^{2+}} = +1.97$ V.

54. How many litres of water must be added to 1 litre an aqueous solution of HCl with a pH of 1 to create an aqueous solution with pH of 2 ?

(1) 0.1 L (2) 0.9 L

(3) 2.0 L (4) 9.0 L

Ans. (4)

Sol. pH = 1 $[H^+] = 10^{-1} = 0.1$ M

pH = 2 $[H^+] = 10^{-2} = 0.01$ M

for dilution of HCl $M_1V_1 = M_2V_2$

$0.1 \times 1 = 0.01 \times V_2$

$V_2 = 10$ lt

Volume of water added = 10 – 1 = 9 litre.

55. The first ionisation potential of Na is 5.1 eV. The value of electron gain enthalpy of Na^+ will be :

(1) –2.55 eV (2) –5.1 eV

(3) –10.2 eV (4) +2.55 eV

Ans. (2)

Sol. $Na \longrightarrow Na^+ + e^-$ I^{st} I.E.

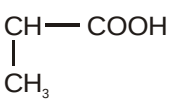
$Na^+ + e^- \longrightarrow Na$ Electron gain enthalpy of Na^+

Because reaction is reverse so then.

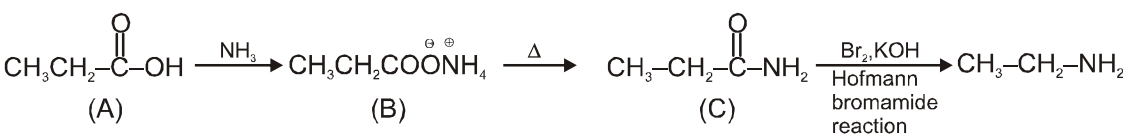
$\Delta H_{eg} = -5.1$ ev.

56. An organic compound A upon reacting with NH_3 gives B. On heating B gives C. C in presence of KOH reacts with Br_2 to given $CH_3CH_2NH_2$. A is :

(1) CH_3COOH (2) $CH_3CH_2CH_2COOH$

(3)  (4) CH_3CH_2COOH

Ans. (4)

Sol. 

57. Stability of the species Li_2, Li_2^- and Li_2^+ increases in the order of :

(1) $Li_2 < Li_2^+ < Li_2^-$ (2) $Li_2^- < Li_2^+ < Li_2$

(3) $Li_2 < Li_2^- < Li_2^+$ (4) $Li_2^- < Li_2 < Li_2^+$

Ans. (2)

Sol. Li_2 $\sigma1s^2$ σ^*1s^2 $\sigma2s^2$ Bond order = 1

Li_2^+ $\sigma1s^2$ σ^*1s^2 $\sigma2s^1$ Bond order = 0.5

Li_2^- $\sigma1s^2$ σ^*1s^2 $\sigma2s^2$ σ^*2s^1 Bond order = 0.5

Stability order $Li_2 > Li_2^+ > Li_2^-$

58. An unknown alochol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism :

(1) secondary alcohol by S_N1 (2) tertiary alcohol by S_N1

(3) secondary alcohol by S_N2 (4) tertiary alcohol by S_N2

Ans. (2)

Sol. The reaction of alcohol with lucas reagent is mostly an S_N1 reaction and the rate of reaction is directly proportional to the carbocation stability formed in the reaction, since 3° R–OH forms 3° carbocation hence it will react fastest.

59. The gas leaked from a storage tank of the Union Carbide plant in Bhopal gas tragedy was :

(1) Methylisocyanate (2) Methylamine

(3) Ammonia (4) Phosgene

Ans. (1)

Sol. Methyl isocyanate $CH_3 – N = C = O$ (MIC gas) (Fact)

60. Experimentally it was found that a metal oxide has formula $M_{0.98}O$. Metal M, present as M^{2+} and M^{3+} in its oxide. Fraction of the metal which exists as M^{3+} would be :

(1) 7.01% (2) 4.08%

(3) 6.05% (4) 5.08%

Ans. (2)

Sol. $M_{0.98}O$

consider one mole of the oxide.

Moles of M = 0.98, Moles of O^{2-} = 1

Let moles of M^{3+} = x

$\Rightarrow$ Moles of M^{2+} = 0.98 – x

$\Rightarrow$ Doing charge balance

$(0.98 – x) \times 2 + 3x – 2 = 0$

$\Rightarrow$ $1.96 – 2x + 3x – 2 = 0$

$\Rightarrow$ $x = 0.04$

$\Rightarrow$ % of $M^{3+} = \frac{0.04}{0.98} \times 100 = 4.08\%$

PART C – MATHEMATICS

61. Distance between two parallel planes $2x + y + 2z = 8$ and $4x + 2y + 4z + 5 = 0$ is

(1) $\frac{3}{2}$ (2) $\frac{5}{2}$ (3) $\frac{7}{2}$ (4) $\frac{9}{2}$

Sol. **(3)**

$2x + y + 2z - 8 = 0$ ...(P_1)

$2x + y + 2z + \frac{5}{2} = 0$ ...(P_2)

Distance between P_1 and P_2 = $\left|\frac{-8-\frac{5}{2}}{\sqrt{2^2+1^2+2^2}}\right| = \frac{7}{2}$

62. At present, a firm is manufacturing 2000 items. It is estimated that the rate of change of production P w.r.t. additional number of workers x is given by $\frac{dP}{dx} = 100 - 12\sqrt{x}$. If the firm employs 25 more workers, then the new level of production of items is

(1) 2500 (2) 3000 (3) 3500 (4) 4500

Sol. **(3)**

$dP = (100 - 12\sqrt{x})dx$

By integrating

$\int dP = \int \left(100 - 12\sqrt{x}\right)dx$

$P = 100x - 8x^{3/2} + C$

When $x = 0$ then $P = 2000$

$\Rightarrow C = 2000$

Now when $x = 25$ then P is

$P = 100 \times 25 - 8 \times (25)^{3/2} + 2000$

$= 2500 - 8 \times 125 + 2000$

$= 4500 - 1000$

$\Rightarrow P = 3500$

63. Let A and B two sets containing 2 elements and 4 elements respectively. The number of subsets of A × B having 3 or more elements is

(1) 256 (2) 220 (3) 219 (4) 211

Sol. **(3)**

$n(A) = 2$

$n(B) = 4$

$n(A \times B) = 8$

${}^8C_3 + {}^8C_4 + \ldots\ldots + {}^8C_8 = 2^8 - {}^8C_0 - {}^8C_1 - {}^8C_2$

$= 256 - 1 - 8 - 28$

$= 219$

64. If the lines $\frac{x-2}{1} = \frac{y-3}{1} = \frac{z-4}{-k}$ and $\frac{x-1}{k} = \frac{y-4}{2} = \frac{z-5}{1}$ are coplanar, then k can have

(1) any value (2) exactly one value (3) exactly two values (4) exactly three values

Sol. **(3)**

$[a - c, b, d] = 0$

$$\begin{vmatrix} 2-1 & 3-4 & 4-5 \\ 1 & 1 & -k \\ k & 2 & 1 \end{vmatrix} = 0$$

$$\begin{vmatrix} 1 & -1 & -1 \\ 1 & 1 & -k \\ k & 2 & 1 \end{vmatrix} = 0$$

$\Rightarrow 1(1 + 2k) + (1 + k^2) - (2 - k) = 0$
$\Rightarrow k^2 + 2k + k = 0$
$\Rightarrow k^2 + 3k = 0$
$\Rightarrow k = 0, -3$

Note : If 0 appears in the denominator, then the correct way of representing the equation of straight line is

$$\frac{x-2}{1} = \frac{y-3}{1};\ z = 4$$

65. If the vectors $\overrightarrow{AB} = 3\hat{i} + 4\hat{k}$ and $\overrightarrow{AC} = 5\hat{i} - 2\hat{j} + 4\hat{k}$ are the sides of a triangle ABC, then the length of the median through A is

(1) $\sqrt{18}$ (2) $\sqrt{72}$ (3) $\sqrt{33}$ (4) $\sqrt{45}$

Sol. **(3)**

$\overrightarrow{AB} + \overrightarrow{BC} + \overrightarrow{CA} = 0$

$\Rightarrow \overrightarrow{BC} = \overrightarrow{AC} - \overrightarrow{AB}$

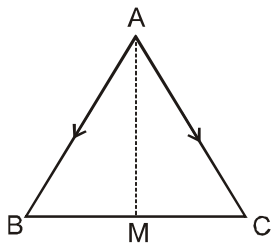
$\Rightarrow \overrightarrow{BM} = \dfrac{\overrightarrow{AC} - \overrightarrow{AB}}{2}$

$\Rightarrow \overrightarrow{AB} + \overrightarrow{BM} + \overrightarrow{MA} = 0$

$\Rightarrow \overrightarrow{AB} + \dfrac{\overrightarrow{AC} - \overrightarrow{AB}}{2} = \overrightarrow{AM}$

$\Rightarrow \overrightarrow{AM} = \dfrac{\overrightarrow{AB} + \overrightarrow{AC}}{2} = 4\hat{i} - \hat{j} + 4\hat{k}$

$\Rightarrow \left|\overrightarrow{AM}\right| = \sqrt{33}$



66. The real number k for which the equation, $2x^3 + 3x + k = 0$ has two distinct real roots in [0, 1]

(1) lies between 1 and 2 (2) lies between 2 and 3

(3) lies between –1 and 0 (4) does not exist.

Sol. **(4)**

$f(x) = 2x^3 + 3x + k$

$f'(x) = 6x^2 + 3 > 0 \quad \forall\ x \in R$

$\Rightarrow$ f(x) is strictly increasing function

$\Rightarrow$ f(x) = 0 has only one real root, so two roots are not possible

67. The sum of first 20 terms of the sequence 0.7, 0.77, 0.777,....., is

(1) $\frac{7}{81}(179 - 10^{-20})$ (2) $\frac{7}{9}(99 - 10^{-20})$ (3) $\frac{7}{81}(179 + 10^{-20})$ (4) $\frac{7}{9}(99 + 10^{-20})$

Sol. **(3)**

$$\frac{7}{10}+\frac{77}{100}+\frac{777}{10^3}+\ldots\ldots\ldots+\text{ up to 20 terms}$$

$$=7\left[\frac{1}{10}+\frac{11}{100}+\frac{111}{10^3}+\ldots..\text{up to 20 terms}\right]$$

$$=\frac{7}{9}\left[\frac{9}{10}+\frac{99}{100}+\frac{999}{1000}+\ldots..\text{up to 20 terms}\right]$$

$$=\frac{7}{9}\left[\left(1-\frac{1}{10}\right)+\left(1-\frac{1}{10^2}\right)+\left(1-\frac{1}{10^3}\right)+\ldots..\text{ up to 20 terms}\right]$$

$$=\frac{7}{9}\left[20-\frac{\frac{1}{10}\left(1-\left(\frac{1}{10}\right)^{20}\right)}{1-\frac{1}{10}}\right]=\frac{7}{9}\left[20-\frac{1}{9}\left(1-\left(\frac{1}{10}\right)^{20}\right)\right]$$

$$=\frac{7}{9}\left[\frac{179}{9}+\frac{1}{9}\left(\frac{1}{10}\right)^{20}\right]=\frac{7}{81}\left[179+(10)^{-20}\right]$$

68. A ray of light along $x+\sqrt{3}y=\sqrt{3}$ gets reflected upon reaching x-axis, the equation of the reflected ray is

(1) $y=x+\sqrt{3}$ (2) $\sqrt{3}y=x-\sqrt{3}$ (3) $y=\sqrt{3}x-\sqrt{3}$ (4) $\sqrt{3}y=x-1$

Sol. **(2)**

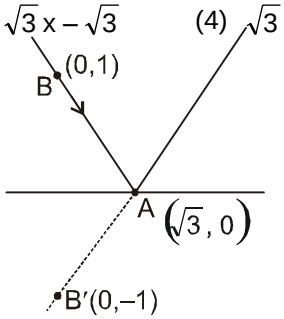
Take any point B(0, 1) on given line

Equation of AB′

$$y-0=\frac{-1-0}{0-\sqrt{3}}\left(x-\sqrt{3}\right)$$

$$-\sqrt{3}y=-x+\sqrt{3}$$

$$x-\sqrt{3}y=\sqrt{3}\quad\Rightarrow\quad\sqrt{3}y=x-\sqrt{3}$$



69. The number of values of k, for which the system of equations :

$(k+1)x+8y=4k$

$kx+(k+3)y=3k-1$

has no solution, is

(1) infinite (2) 1 (3) 2 (4) 3

Sol. **(2)**

$$\frac{k+1}{k}=\frac{8}{k+3}\neq\frac{4k}{3k-1}$$

$k^2+4k+3=8k$

$k^2-4k+3=0$

$k=1,3$

If $k=1$

then $\frac{8}{1+3}\neq\frac{4.1}{2}$ False

And If $k=3$

then $\frac{8}{6}\neq\frac{4.3}{9-1}$ True

therefore $k=3$

Hence only one value of k.

70. If the equations $x^2 + 2x + 3 = 0$ and $ax^2 + bx + c = 0$, $a,b,c \in R$, have a common root, then a : b : c is

(1) 1 : 2 : 3 (2) 3 : 2 : 1 (3) 1 : 3 : 2 (4) 3 : 1 : 2

Sol. **(1)**

$x^2 + 2x + 3 = 0$ ...(i)

$ax^2 + bx + c = 0$ ...(ii)

Since equation (i) has imaginary roots

So equation (ii) will also have both roots same as (i). Thus

$$\frac{a}{1} = \frac{b}{2} = \frac{c}{3} \quad \Rightarrow \quad a = \lambda,\ b = 2\lambda,\ c = 3\lambda$$

Hence 1 : 2 : 3

71. The circle passing through (1, –2) and touching the axis of x at (3, 0) also passes through the point

(1) (–5, 2) (2) (2, – 5) (3) (5, – 2) (4) (–2, 5)

Sol. **(3)**

Let the equation of circle be

$(x – 3)^2 + (y – 0)^2 + \lambda y = 0$

As it passes through (1, – 2)

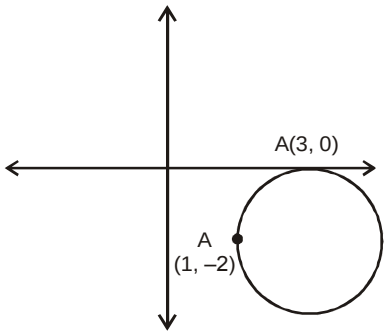
$\therefore \quad (1 – 3)^2 + (– 2)^2 + \lambda(–2) = 0$

$\Rightarrow \quad \lambda = 4$

$\therefore$ equation of circle is

$(x – 3)^2 + y^2 – 8 = 0$

so (5, – 2) satisfies equation of circle



72. If x, y, z are in A.P. and $\tan^{-1}x$, $\tan^{-1}y$ and $\tan^{-1}z$ are also in A.P., then

(1) x = y = z (2) 2x = 3y = 6z (3) 6x = 3y = 2z (4) 6x = 4y = 3z

Sol. **(1)**

$2y = x + z$

$2\tan^{-1} y = \tan^{-1} x + \tan^{-1} (z)$

$$\tan^{-1}\left(\frac{2y}{1-y^2}\right) = \tan^{-1}\left(\frac{x+z}{1-xz}\right)$$

$$\frac{x+z}{1-y^2} = \frac{x+z}{1-xz}$$

$\Rightarrow \quad y^2 = xz$ or $x + z = 0 \quad \Rightarrow \quad x = y = z$

73. Consider

Statement-I : $(p \wedge \sim q) \wedge (\sim p \wedge q)$ is a fallacy.

Statement-II : $(p \to q) \leftrightarrow (\sim q \to \sim p)$ is a tautology.

(1) Statement-I is true; Statement-II is true; Statement-II is a correct explanation for Statement-I.

(2) Statement-I is true; Statement-II is true; Statement-II is **not** a correct explanation for Statement-I.

(3) Statement-I is true; Statement-II is false.

(4) Statement-I is false; Statement-II is true.

Sol. **(2)**

Statement-II : $(p \to q) \leftrightarrow (\sim q \to \sim p)$

$\equiv (p \to q) \leftrightarrow (p \to q)$

which is always true

so statement -II is true

Statement-I : $(p \wedge \sim q) \wedge (\sim p \wedge q)$

$= p \wedge \sim q \wedge \sim p \wedge q$

$= p \wedge \sim p \wedge \sim q \wedge q$

$= f \wedge f$

$= f$

so statement -I is true

Alternate

Statement-II : $(p \to q) \leftrightarrow (\sim q \to \sim p)$

$\sim q \to \sim p$ is contrapositive

of $p \to q$ hence $(p \to q) \leftrightarrow (p \to q)$

will be a tautology

statement-II $(p \wedge \sim q) \wedge (\sim p \wedge q)$

p	q	P ^ ~ q	~p ^ q	(p ^ ~ q) ^ (~ p ^ q)
T	T	F	F	F
T	F	T	F	F
F	T	F	T	F
F	F	F	F	F

$\therefore$ It is a fallacy

74. If $\int f(x)\,dx = \psi(x)$, then $\int x^5 f(x^3)\,dx$ is equal to

(1) $\frac{1}{3}\left[x^3\psi(x^3) - \int x^2\psi(x^3)\,dx\right] + C$

(2) $\frac{1}{3}x^3\psi(x^3) - 3\int x^3\psi(x^3)\,dx + C$

(3) $\frac{1}{3}x^3\psi(x^3) - \int x^2\psi(x^3)\,dx + C$

(4) $\frac{1}{3}\left[x^3\psi(x^3) - \int x^3\psi(x^3)\,dx\right] + C$

Sol. **(3)**

$\int f(x)dx = \psi(x)$

$I = \int x^5 f(x^3)dx$

put $x^3 = t \quad \Rightarrow \quad x^2dx = \frac{dt}{3}$

$= \frac{1}{3}\int tf(t)dt$

$= \frac{1}{3}\left[t\psi(t) - \int \psi(t)dt\right]$

$= \frac{1}{3}\left[x^3\psi(x^3) - 3\int x^2\psi(x^3)dx\right] + c$

$= \frac{1}{3}x^3\psi(x^3) - \int x^2\psi(x^3)dx + c$

75. $\lim\limits_{x\to 0} \frac{(1-\cos 2x)(3+\cos x)}{x\tan 4x}$ is equal to

(1) $-\frac{1}{4}$ (2) $\frac{1}{2}$ (3) 1 (4) 2

Sol. **(4)**

$I = \lim\limits_{x\to 0} \frac{(1-\cos 2x)}{x^2}\frac{(3+\cos x)}{1}.\frac{x}{\tan 4x}$

$= \lim\limits_{x\to 0} \frac{2\sin^2 x}{x^2}.\frac{3+\cos x}{1}.\frac{x}{\tan 4x}$

$= 2.4.\frac{1}{4} = 2$

76. Statement-I : The value of the integral $\int_{\pi/6}^{\pi/3} \frac{dx}{1+\sqrt{\tan x}}$ is equal to $\pi/6$.

Statement-II : $\int_a^b f(x)dx = \int_a^b f(a+b-x)\,dx$.

(1) Statement-I is true; Statement-II is true; Statement-II is a **correct** explanation for Statement-I.
(2) Statement-I is true; Statement-II is true; Statement-II is **not** a correct explanation for Statement-I.
(3) Statement-I is true; Statement-II is false.
(4) Statement-I is false; Statement-II is true.

Sol. **(4)**

$$I = \int_{\pi/6}^{\pi/3} \frac{dx}{1+\sqrt{\tan x}}$$

$$= \int_{\pi/6}^{\pi/3} \frac{dx}{1+\sqrt{\tan\left(\frac{\pi}{2}-x\right)}}$$

$$= \int_{\pi/6}^{\pi/3} \frac{\sqrt{\tan x}\,dx}{1+\sqrt{\tan x}}$$

$$= \int_{\pi/6}^{\pi/3} \frac{\sqrt{\tan x}\,dx}{1+\sqrt{\tan x}}$$

$$\Rightarrow \quad 2I = \int_{\pi/6}^{\pi/3} dx$$

$$\Rightarrow \quad I = \frac{1}{2}\left[\frac{\pi}{3}-\frac{\pi}{6}\right] = \frac{\pi}{12}, \quad \text{statement -1 is false}$$

$\int_a^b f(x)dx = \int_a^b f(a+b-x)dx$ it is property

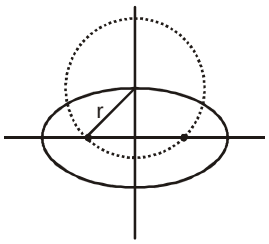
77. The equation of the circle passing through the foci of the ellipse $\frac{x^2}{16}+\frac{y^2}{9} = 1$, and having centre at (0, 3) is

(1) $x^2 + y^2 - 6y - 7 = 0$ (2) $x^2 + y^2 - 6y + 7 = 0$
(3) $x^2 + y^2 - 6y - 5 = 0$ (4) $x^2 + y^2 - 6y + 5 = 0$

Sol. **(1)**

$a = 4, b = 3, e = \sqrt{1-\frac{9}{16}} \quad \Rightarrow \quad \frac{\sqrt{7}}{4}$

Focii is $(\pm ae, 0) \quad \Rightarrow \quad (\pm\sqrt{7}, 0)$



$r = \sqrt{(ae)^2+b^2}$

$\sqrt{7+9}$

$= 4$

Now equation of circle is $(x-0)^2 + (y-3)^2 = 16$

$x^2 + y^2 - 6y - 7 = 0$

78. A multiple choice examination has 5 questions. Each question has three alternative answers of which exactly one is correct. The probability that a student will get 4 or more correct answers just by guessing is :

(1) $\frac{17}{3^5}$ (2) $\frac{13}{3^5}$ (3) $\frac{11}{3^5}$ (4) $\frac{10}{3^5}$

Sol. **(3)**

$p = \frac{1}{3}$, $q = \frac{2}{3}$

$^5C_4 \left(\frac{1}{3}\right)^4 \cdot \frac{2}{3} + {}^5C_5 \left(\frac{1}{3}\right)^5$

$= 5. \frac{2}{3^5} + \frac{1}{3^5} = \frac{11}{3^5}$

79. The x-coordinate of the incentre of the triangle that has the coordinates of mid points of its sides as (0, 1) (1, 1) and (1, 0) is :

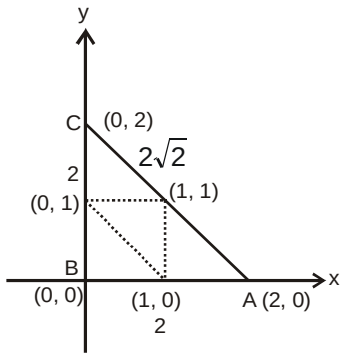
(1) $2 + \sqrt{2}$ (2) $2 - \sqrt{2}$ (3) $1 + \sqrt{2}$ (4) $1 - \sqrt{2}$

Sol. **(2)**

x - coordinate of incentre $= \frac{2\times 0 + 2\sqrt{2}.0 + 2.2}{2+2+2\sqrt{2}}$

$= \frac{2}{2+\sqrt{2}}$

$= 2 - \sqrt{2}$



80. The term independent of x in expansion of $\left(\frac{x+1}{x^{2/3}-x^{1/3}+1} - \frac{x-1}{x-x^{1/2}}\right)^{10}$ is :

(1) 4 (2) 120 (3) 210 (4) 310

Sol. **(3)**

$\left((x^{1/3}+1) - \left(\frac{\sqrt{x}+1}{\sqrt{x}}\right)\right)^{10}$

$(x^{1/3} - x^{-1/2})^{10}$

$T_{r+1} = {}^{10}c_r (x^{1/3})^{10-r} (-x^{-1/2})^r$

$\frac{10-r}{3} - \frac{r}{2} = 0 \Rightarrow 20 - 2r - 3r = 0$

$\Rightarrow r = 4$

$T_5 = {}^{10}C_4 = \frac{10\times 9\times 8\times 7}{4\times 3\times 2\times 1} = 210$

81. The area (in square units) bounded by the curves $y = \sqrt{x}$, $2y - x + 3 = 0$, x-axis, and lying in the first quadrant is :

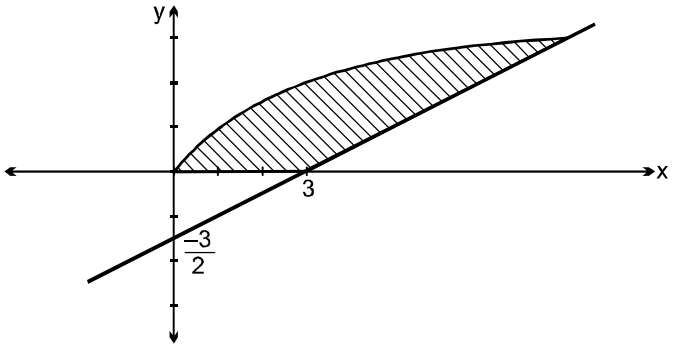
(1) 9 (2) 36 (3) 18 (4) $\frac{27}{4}$

Sol. **(1)**

$y = \sqrt{x}$ .....(1)

and $2y - x + 3 = 0$ .....(2)

On solving both $y = -1, 3$



Required area = $\int_0^3 \left\{(2y+3) - y^2\right\} dy$

$= y^2 + 3y - \frac{y^3}{3}\Big|_0^3$

$= 9 + 9 - 9$

$= 9.$

82. Let T_n be the number of all possible triangles formed by joining vertices of an n-sided regular polygon. If $T_{n+1} - T_n = 10$, then the value of n is :

(1) 7 (2) 5 (3) 10 (4) 8

Sol. **(2)**

$T_n = {}^nC_3$

$T_{n+1} = {}^{n+1}C_3$

$T_{n+1} - T_n = {}^{n+1}C_3 - {}^nC_3$

$\Rightarrow \quad {}^nC_2 = 10$

$\Rightarrow \quad n = 5.$

83. If z is a complex number of unit modulus and argument θ, then arg $\left(\frac{1+z}{1+\bar{z}}\right)$ equals :

(1) $-\theta$ (2) $\frac{\pi}{2} - \theta$ (3) θ (4) $\pi - \theta$

Sol. **(3)**

$|z| = 1$, arg $z = \theta$ $\quad z = e^{i\theta}$

$\bar{z} = \frac{1}{z}$

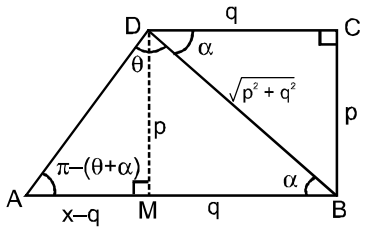
arg $\left(\frac{1+z}{1+\frac{1}{z}}\right)$ = arg $(z) = \theta$.

84. ABCD is a trapezium such that AB and CD are parallel and BC $\perp$ CD. If $\angle ADB = \theta$, BC = p and CD = q, then AB is equal to :

(1) $\frac{(p^2+q^2)\sin\theta}{p\cos\theta + q\sin\theta}$ (2) $\frac{p^2 + q^2\cos\theta}{p\cos\theta + q\sin\theta}$ (3) $\frac{p^2+q^2}{p^2\cos\theta + q^2\sin\theta}$ (4) $\frac{(p^2+q^2)\sin\theta}{(p\cos\theta + q\sin\theta)^2}$

Sol. **(1)**

Let AB = x



$\tan(\pi-\theta-\alpha)=\frac{p}{x-q} \Rightarrow \tan(\theta+\alpha)=\frac{p}{q-x}$

$\Rightarrow$ $q-x=p\cot(\theta+\alpha)$

$\Rightarrow$ $x=q-p\cot(\theta+\alpha)$

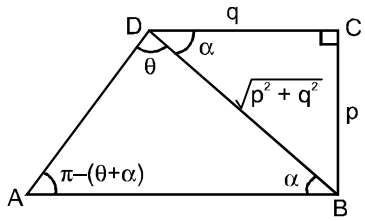
$$=q-p\left(\frac{\cot\theta\cot\alpha-1}{\cot\alpha+\cot\theta}\right)$$

$$=q-p\left(\frac{\frac{q}{p}\cot\theta-1}{\frac{q}{p}+\cot\theta}\right)=q-p\left(\frac{q\cot\theta-p}{q+p\cot\theta}\right)=q-p\left(\frac{q\cos\theta-p\sin\theta}{q\sin\theta+p\cos\theta}\right)$$

$\Rightarrow$ $x=\frac{q^2\sin\theta+pq\cos\theta-pq\cos\theta+p^2\sin\theta}{p\cos\theta+q\sin\theta} \Rightarrow AB=\frac{(p^2+q^2)\sin\theta}{p\cos\theta+q\sin\theta}$.

Alternative

From Sine Rule



$$\frac{AB}{\sin\theta}=\frac{\sqrt{p^2+q^2}}{\sin(\pi-(\theta+\alpha))}$$

$$AB=\frac{\sqrt{p^2+q^2}\sin\theta}{\sin\theta\cos\alpha+\cos\theta\sin\alpha}$$

$$=\frac{(p^2+q^2)\sin\theta}{q\sin\theta+p\cos\theta} \qquad \left(\because \cos\alpha=\frac{q}{\sqrt{p^2+q^2}}\right)$$

$$=\frac{(p^2+q^2)\sin\theta}{p\cos\theta+q\sin\theta}.$$

85. If $P=\begin{bmatrix}1 & \alpha & 3\\ 1 & 3 & 3\\ 2 & 4 & 4\end{bmatrix}$ is the adjoint of a 3 × 3 matrix A and |A| = 4, then α is equal to :

(1) 4 (2) 11 (3) 5 (4) 0

Sol. **(2)**

$|P| = 1(12-12)-\alpha(4-6)+3(4-6)$

$= 2\alpha - 6$

$|P| = |A|^2 = 16$

$2\alpha - 6 = 16$

$\alpha = 11.$

86. The intercepts on x-axis made by tangents to the curve, $y=\int_0^x |t|dt, x \in R$, which are parallel to the line y = 2x, are equal to :

(1) ±1 (2) ±2 (3) ±3 (4) ±4

Sol. **(1)**

$\frac{dy}{dx} = |x| = 2$

$x = \pm 2$

points $y=\int_0^{\pm 2}|t|dt = \pm 2$

$\therefore$ equation of tangent is

$y - 2 = 2(x-2)$ or $y + 2 = 2(x+2)$ $\Rightarrow$ x - intercept = ± 1.

87. Given : A circle, $2x^2 + 2y^2 = 5$ and a parabola, $y^2 = 4\sqrt{5}\,x$.

Statement-I : An equation of a common tangent to these curves is $y = x + \sqrt{5}$.

Statement-II : If the line, $y = mx + \frac{\sqrt{5}}{m}$ $(m \neq 0)$ is their common tangent, then m satisfies $m^4 - 3m^2 + 2 = 0$.

(1) Statement-I is true; Statement-II is true; Statement-II is a **correct** explanation for Statement-I.
(2) Statement-I is true; Statement-II is true; Statement-II is **not** a correct explanation for Statement-I.
(3) Statement-I is true; Statement-II is false.
(4) Statement-I is false; Statement-II is true.

Sol. **(2)**

Let common tangent

$$y = mx + \frac{\sqrt{5}}{m}$$

$$\frac{\frac{\sqrt{5}}{m}}{\sqrt{1+m^2}} = \sqrt{\frac{5}{2}}$$

$$m\sqrt{1+m^2} = \sqrt{2}$$
$$m^2(1 + m^2) = 2$$
$$m^4 + m^2 - 2 = 0$$
$$(m^2 + 2)(m^2 - 1) = 0$$
$$m = \pm 1$$

$y = \pm (x + \sqrt{5})$, both statements are correct as $m = \pm 1$ satisfies the given equation of statement-2.

88. If $y = \sec(\tan^{-1}x)$, then $\frac{dy}{dx}$ at $x = 1$ is equal to :

(1) $\frac{1}{\sqrt{2}}$ (2) $\frac{1}{2}$ (3) 1 (4) $\sqrt{2}$

Sol. **(1)**

$y = \sec(\tan^{-1} x)$

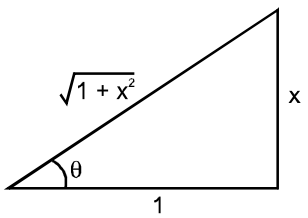
Let $\tan^{-1} x = \theta$

$x = \tan\theta$

$y = \sec\theta$

$$y = \sqrt{1+x^2}$$

$$\frac{dy}{dx} = \frac{1}{2\sqrt{1+x^2}} \cdot 2x$$



at $x = 1$

$$\frac{dy}{dx} = \frac{1}{\sqrt{2}}.$$

89. The expression $\frac{\tan A}{1-\cot A} + \frac{\cot A}{1-\tan A}$ can be written as :

(1) sinA cosA + 1 (2) secA cosecA + 1
(3) tanA + cotA (4) secA + cosecA

Sol. **(2)**

Given expression

$$= \frac{\sin A}{\cos A} \times \frac{\sin A}{\sin A - \cos A} + \frac{\cos A}{\sin A} \times \frac{\cos A}{\cos A - \sin A}$$

$$= \frac{1}{\sin A - \cos A} \left\{ \frac{\sin^3 A - \cos^3 A}{\cos A \sin A} \right\}$$

$$= \frac{\sin^2 A + \sin A \cos A + \cos^2 A}{\sin A \cos A} = 1 + \sec A \operatorname{cosec} A$$

90. All the students of a class performed poorly in Mathematics. The teacher decided to give grace marks of 10 to each of the students. Which of the following statistical measures will not change even after the grace marks were given ?

(1) mean (2) median (3) mode (4) variance

Sol. **(4)**

If initially all marks were x_i then $\sigma_1^2 = \frac{\sum (x_i - \bar{x})^2}{N}$

Now each is increased by 10

$$\sigma_2^2 = \frac{\sum [(x_i + 10) - (\bar{x} + 10)]^2}{N} = \sigma_1^2$$

So variance will not change whereas mean, median and mode will increase by 10.