

**Class: X**  
**Subject: chemistry**  
**Topic: Acids, Bases and Salts**  
**No. of Questions: 20**  
**Duration: 60 Min**  
**Maximum Marks: 60**

**1: Matter has \_\_\_\_\_.**

1. no mass but occupies space
2. mass but occupies no space
3. mass and occupies space
4. no mass and occupies no space

**Answer: 3**

**Explanation: Matter is that** Physical substance in general, as distinct from mind and spirit; (in physics) that which occupies space and possesses rest mass, especially as distinct from energy

**2: The gaseous form of water is called as \_\_\_\_\_.**

1. water gas
2. water vapour
3. fog
4. snow

**Answer: 2**

**Explanation: Water vapor, or water vapour or aqueous vapor, is the gaseous phase of water.** It is one state of **water** within the hydrosphere. **Water vapor** can be produced from the evaporation or boiling of liquid **water** or from the sublimation of ice. Unlike other forms of **water, water vapor** is invisible.

**3: The state of matter with only one free surface is \_\_\_\_\_.**

1. liquid
2. gas
3. solid
4. plasma

**Answer:** 1

**Explanation:** A **liquid** is a nearly incompressible fluid that conforms to the shape of its container but retains a (nearly) constant volume independent of pressure. As such, it is one of the four fundamental states of matter (the others being solid, gas, and plasma), and is the only state with a definite volume but no fixed shape.

**4: In gases, particles vibrate \_\_\_\_\_.**

1. about their mean position
2. about a vertical axis
3. about a horizontal axis
4. in any direction

**Answer:** 4

**Explanation:** The particles of gas, either atoms or molecules, have too much energy to remain attached to one other. They move by translation, rotation and vibration, but in this case the translational motion is the most important. The particles are on average very far apart and collide incessantly many times a second. Because of the distance between them it is assumed that the forces of attraction between the particles are negligible.

**5: Mixing of gases is called \_\_\_\_\_.**

1. diffusion
2. effusion
3. filtration
4. sedimentation

**Answer:** 1

**Explanation:** Diffusion involves mixing of gases

**6: A property not possessed by a fluid is \_\_\_\_\_.**

1. it can flow
2. it has mass
3. it has a definite shape
4. can be perceived by our senses

**Answer:** 3

**Explanation:** Fluids don't have definite shape. They take the shape of the Container

**7: Which among the following is a solid at room temperature?**

1. Nitrogen
2. Potassium permanganate
3. Bromine
4. Helium

**Answer:** 2

**8: Which of these is a characteristic property of gases?**

1. Gases are not at all rigid.
2. Gases are not compressible.
3. Gases have particles in fixed positions.
4. Gases have high density.

**Answer:** 1

**9: A solid has \_\_\_\_\_.**

1. maximum intermolecular space
2. definite mass but no definite volume
3. very high compressibility
4. maximum intermolecular force of attraction

**Answer:** 4

**10: Conversion of gas to liquid is called \_\_\_\_\_.**

- 1 condensation
- 2 sublimation
- 3 vapourisation
- 4 solidification

**Answer:** 1

**Explanation:** **Condensation** is the change of water from its gaseous form (water vapor) into liquid water. **Condensation** generally occurs in the atmosphere when warm air rises, cools and loses its capacity to hold water vapor. As a result, excess water vapor **condenses** to form cloud droplets.

**11: The process by which wet clothes dry up is called \_\_\_\_\_.**

1. evaporation
2. boiling
3. condensation
4. solidification

**Answer:** 1

**Explanation:** **Evaporation** is a type of vaporization of a liquid that occurs from the surface of a liquid into a gaseous phase that is not saturated with the **evaporating** substance. The other type of vaporization is boiling, which is characterized by bubbles of saturated vapor forming in the liquid phase

**12: A solid that sublimates on heating is \_\_\_\_\_.**

1. sodium chloride
2. copper sulphate
3. lead sulphate
4. ammonium chloride

**Answer:** 4

**13: Matter changes from one state to another with change in \_\_\_\_\_.**

1. density
2. temperature
3. volume
4. height

**Answer:** 2

**Explanation:** With change in temperature matter changes its phase

**14: The process of a solid changing into liquid is called \_\_\_\_\_.**

1. liquefaction
2. melting
3. freezing
4. solidification

**Answer:** 2

**Explanation:** Freezing is a phase transition in which a liquid turn into a solid when its temperature is lowered to its freezing point

**15: The freezing point of pure water is \_\_\_\_\_.**

1. 100°C
2. 0°C
3. 5°C
4. 78.3°C

**Answer:** 2

**16: The force between particles of matter is called as \_\_\_\_\_.**

1. cohesive force
2. adhesive force
3. kinetic energy
4. thermal energy

**Answer:** 1

**Explanation:** Cohesive force is the force of attraction between the molecules of the same substance. Cohesive force is stronger in solids than in liquid.  
Cohesive force is stronger in liquids than in gas.

**17: Particles of matter are \_\_\_\_\_.**

1. stationary
2. vibrating in one position
3. in continuous motion
4. rotating about an axis

**Answer:** 3

**18: The temperature at which vapour changes into liquid is called \_\_\_\_\_.**

1. freezing point
2. melting point
3. boiling point
4. liquefaction point

**Answer:** 4

**Explanation:** **Liquefaction**, sometimes **liquification**, refers to any process which either generates a liquid from a solid or a gas, or generates a non-liquid phase which behaves in accordance with fluid dynamics

**19: Temperature is a measure of \_\_\_\_\_.**

1. total kinetic energy of molecules
2. total potential energy of molecules
3. average potential energy of molecules
4. average kinetic energy of molecules

**Answer:** 1

**Explanation:** Temperature is a measure of the average heat or thermal energy of the particles in a substance. Since it is an average measurement, it does not depend on the number of particles in an object. In that sense it does not depend on the size of it. For example, the temperature of a small cup of boiling water is the same as the temperature of a large pot of boiling water. Even if the large pot is much bigger than the cup and has millions and millions more water molecules.

**20: For any substance the temperature remains same during the change of state due to \_\_\_\_\_.**

1. loss of heat
2. latent heat
3. less supply of heat
4. lattice energy

**Answer:** 2

**Explanation:** **Latent heat** is energy released or absorbed, by a body or a thermodynamic system, during a constant-temperature process. An example is a state of matter change, meaning a phase transition, such as ice melting or water boiling.