

Class: X  
Subject: Chemistry  
Topic: Acids, bases and salts  
No. of Questions: 20

- Q1. Why should curd and sour substances not be kept in brass and copper vessels?
- Q2. How would you distinguish between baking powder and washing soda by heating?
- Q3. An element common to all acid is
- Chlorine
  - Nitrogen
  - Oxygen
  - Hydrogen
- Q4. Why does dry HCl gas not change the colour of the dry litmus paper?
- Q5. Name the products formed from the chloro-alkali process.
- Q6. Why do HCl, HNO<sub>3</sub>, etc, show acidic characters in aqueous solutions while solutions of compounds like alcohol and glucose do not show acidic character?
- Q7. Why aqueous solution of sodium carbonate is basic in nature?
- Q8. On passing excess of CO<sub>2</sub> gas in an aqueous solution of calcium carbonate. Milkiness of the solution
- persists
  - fades
  - deepens
  - disappears
- Q9. What is a neutralisation reaction ? Give two examples.
- Q10. What happens when a solution of sodium hydrogen carbonate is heated ? Write equation of the reaction involved.

- Q11. Why does an aqueous solution of an acid conduct electricity?
- Q12. Write the composition of baking powder. What will happen if tartaric acid is not added to it?
- Q13. What is deliquescent and hygroscopic substance?
- Q14. How is plaster of paris chemically different from gypsum?
- Q15. What happens to most of the molecules of a strong acid when the acid is mixed with water?
- Q16. What are hydrated salts? Give an example.
- Q17. Tartaric acid is the constituent of
- A. Bleaching powder
  - B. Baking powder
  - C. Washing powder
  - D. Plaster of paris
- Q18. Do basic solutions also have  $H^+$  (aq) ions. If yes, then why are these basic?
- Q19. Name the organic acid present in tomato
- A. Tartaric Acid
  - B. Malic Acid
  - C. Lactic Acid
  - D. Oxalic Acid
- Q20. Write the balanced equation in molecular form illustrating the complete neutralization of  $Al(OH)_3$  with  $H_2SO_4$ .