

Class: 11
Subject: Chemistry
Topic: Hydrogen
No. of Questions: 20

- Q1. Discuss the structure of water.
- Q2. Write the uses of dihydrogen?
- Q3. Explain the amphoteric nature of water.
- Q4. Comment on any one laboratory method of preparation of dihydrogen?
- Q5. How hydrogen resembles alkali metals and halogens?
- Q6. Why H^+ ion does not exist freely?
- Q7. Give the industrial method of preparation of hydrogen peroxide.
- Q8. Give an example of coordinated water molecule.
- Q9. Why does hydrogen occur in a diatomic form rather than in a monoatomic form under normal conditions?
- Q10. What are ionic hydrides?
- Q11. What happens when dihydrogen reacts with dinitrogen?
- Q12. How hydrogen differs from alkali metals?
- Q13. Give an example of hydrogen bonded water molecule.

- Q14. Give an example of interstitial water molecule.
- Q15. Discuss the consequences of high enthalpy of H-H bond in terms of chemical reactivity of dihydrogen.
- Q16. What is the reaction of dihydrogen with metals?
- Q17. Give four uses of hydrogen peroxide.
- Q18. How do you expect the metallic hydrides to be useful for hydrogen storage? Explain.
- Q19. H_2O_2 can be used as:
a) Bleaching agent
b) Oxidizing agent
c) Antiseptic
d) All the above Solution
- Q20. Water has maximum density at:
a) 0°C
b) 1°C
c) 4°C
d) 4K